

15. Test your understanding of isotopes by examining the relationships between the pairs of atoms listed below:

Atom 1	Atom 2	Relationship between atom 1 and atom 2
$^{12}_6 C$	$^{13}_6 C$	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element
Carbon-12	$^{12}_6 C$	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element
Argon-40	Argon-41	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element
$^{11}_5 B$	Boron-10	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element
An atom with 13 protons and 13 neutrons	An atom with 14 protons and 13 neutrons	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element

EXERCISES

16. The periodic table has a great deal of information about every atom. Using your periodic table, answer the following questions:

- What is the atomic number of chlorine (Cl)? 17
- What is the atomic number of tungsten (W)? 74
- How many protons are there in any Cl atom? 17
- How many protons are there in any Te atom? 52
- Can you tell from the periodic table exactly how many neutrons are in an atom?

No, the mass number is an average mass of all of the isotopes of that element. Since isotopes are based on the number of neutrons, the mass of each isotope of an element will be a little different.

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**National Cattlemen's Beef Association
(U.S.)**

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Assessment that Informs Practice ,2000 ENC Focus ,2000 **Resources in Education** ,1995 *Instructor's Resource Manual* Joanne McNergney,Robert McNergney,2004-02 **Glencoe Mathematics** William Collins,1999

California Science ,2008 Science stimulates curiosity and student inquiry integrates powerful support for reading and science literacy reaches all learners through numerous components and strategies for differentiated instruction reinforces learning through exciting visuals and electronic components and makes teaching science easy with a variety of teacher resources *Glencoe Science* McGraw-Hill Staff,2001-08 *Te Math Advantage 98 Gr 6 Vol 1* Harcourt Brace Publishing,1997-10-10 *Glencoe Science* ,2002 Proceedings of the High School Conference University of Illinois. High school visitor,1919 **Proceedings of the High School Conference of November 1910-November 1931** Horace Adelbert Hollister,1921 *Proceedings of the High School Conference of November 1910-November 1931* ,1921 **Our World's Story** ,1997 **Parliamentary Debates (Hansard)**. Great Britain. Parliament. House of Commons,2009

School ,1913 **Unit Instruction Sheets and Individual Instruction in Vocational Classes** Robert Henry Rodgers,Oakley Furney,1922 *Current Index to Journals in Education* ,1999-07 **Christian Science Sentinel** ,1898

Invitations to Literacy James David Cooper,1997 The NCBA Research Annual Report National Cattlemen's Beef Association (U.S.),2003

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