

A student investigates the evaporation of water. He pours 100cm^3 of water into measuring cylinder A and 100cm^3 of water into measuring cylinder B. Measuring cylinder A is kept at 40°C and B is kept at 80°C in the same part of the laboratory. Fig. 2.1 shows the two measuring cylinders after 3 days.



Fig. 2.1

(a) (i) Using ideas about molecules, explain why more water evaporates when the water is kept at 80°C .

¹² See, for example, the discussion of the 1992 Constitutional Convention in the *Constitutional Convention of 1992: The Final Report of the Constitutional Commission* (1993).

[2]

(ii) Apart from an increase in temperature, state one change that causes water to evaporate faster.

[1]

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(b) The specific latent heat of vaporisation of water is 2300 J/g and the density of water is 1.0 g/cm³. During the three days, the water level in B drops from the 100 cm³ mark to the level shown in Fig. 2.1. Calculate the energy used to evaporate water from B during the three days.

energy = [2]

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