

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations) (opposite charges).



Covalent ionic and polar covalent bonds

- Covalent bonds are classified as ionic (polar) and non-polar.
- Ionic bonds are formed by the transfer of electrons.
- Covalent bonds are formed by sharing of electrons between two atoms.
- Non-polar covalent bonds consist of two equal and identical atoms.



Non-polar covalent bonds are stronger than ionic.

Four types of bonds

Dependence of bond length



Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity between two atoms, polar bonds can be polar and nonpolar.

As a rough measure:

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond. (Nonpolar examples are the C-C and C-H bonds.)
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond. Nonpolar examples are the carbon-carbon in N2, O2, F and other halogens.
- A difference of 1.7 or higher one of both the atoms can have a larger share and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Nucleophiles



Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Direction may be indicated by arrows on the dipole moment.

- The dipole moment is shown as a vector pointing from the less electronegative toward the more electronegative atom.
- The vector is proportional to the partial negative charge and the distance between a partial positive charge and a partial negative charge.
- The dipole plot indicates the charge difference.
- The dipole moment indicates the charge and area.

Molecular Dipole

To measure the molecular polarity, the total sum of the dipole is used.

The polarity of the molecule is equal to the sum of individual bond dipole moments. If dipole moments are canceled, it just means a neutral dipole.



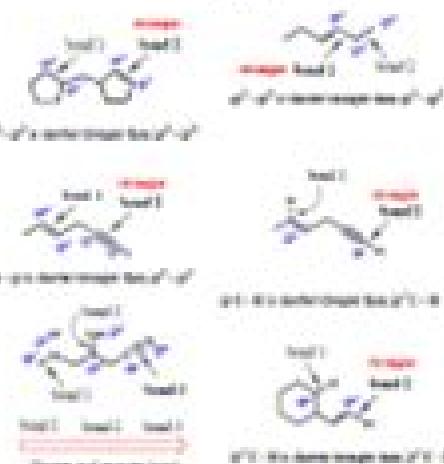
Sum of vector of the bond dipole for the molecule.

Average values of C-C and C-H bond depending on electronegativity

C-C Bond	$\mu_{\text{C-C-C}} = 0.12 \text{ D}$	$\mu_{\text{C-C-H}} = 0.16 \text{ D}$
Hybridization	sp^3 (0.12)	sp^3 (0.16)
Length	1.54 Å	1.10 Å
Strength	0.16 kcal/mol	0.17 kcal/mol
Hybridization	sp^2	sp^2
Hybridization	sp	sp
Length	1.39 Å	1.09 Å
Strength	0.19 kcal/mol	0.20 kcal/mol

Dipole moment and bond length

Example



Bond Length and Bond Strength

Relative bond length



Length: $\text{C-C} > \text{C-O} > \text{C-Cl} > \text{C-F}$

Strength: $\text{C-C} < \text{C-O} < \text{C-Cl} < \text{C-F}$

Partial bond length: $\text{C-C} > \text{C-O} > \text{C-Cl} > \text{C-F}$

Relative bond length: $\text{C-C} > \text{C-O} > \text{C-Cl} > \text{C-F}$

Relative bond strength: $\text{C-C} < \text{C-O} < \text{C-Cl} < \text{C-F}$

Relative bond length



Relative bond length

Length: $\text{C-C} > \text{C-O} > \text{C-Cl} > \text{C-F}$

Strength: $\text{C-C} < \text{C-O} < \text{C-Cl} < \text{C-F}$

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Michal Rosen-Zvi

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