

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds

1. Ionic bond
2. Covalent bond
3. Hydrogen bond
4. Metallic bond



The Behavior of the bond



Strong and Weak bond



Distance between atoms



How are the four different definitions of bonds are

Ionic bond



Covalent bond



Hydrogen bond



Metallic bond



- Ionic bonds are formed by the transfer of electron from one atom to another (non-polar to polar).
- Covalent bonds are formed by sharing of electron between two atoms.
- Hydrogen bonds are formed by sharing of electron between two atoms.
- Metallic bonds are formed by sharing of electron between two atoms.

Polar and Nonpolar Covalent Bonds

Depending on the difference
in the electronegativity
value, polar bonds can
be polar and nonpolar.



- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar or non-bonded bond. (Nonpolar example: H_2S and H_2O bonds).
- If the difference in electronegativity is between 0.5 and 1.7, the bond is a polar covalent bond. Nonpolar example: C_2H_6 for the carbon in C_2H_6 , O_2 and other halogens.
- A difference of 1.7 or higher one of both the atoms can have a larger share and a polar bond is formed.

A Summary of Ring-opening Reactions of Epoxydes with Strong and Weak Protophiles



Preparation and Reactions

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).

1. Nonpolar covalent bond
2. Polar covalent bond
3. Ionic bond



- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector is longer the more partial negative charge and the less electronegative atom has a partial positive charge.
- The dipole moment indicates the difference in electronegativity.
- The dipole moment indicates the difference in electronegativity.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the dipole moments are canceled, it is called a neutral dipole.



Sum of vector of the bond dipole moment for the molecule.

Bond Length and Bond Strength

Relative bond length

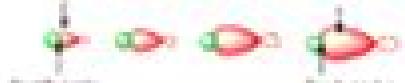


Length: Bond 1 > Bond 2 > Bond 3 > Bond 4

Strength: Bond 1 < Bond 2 < Bond 3 < Bond 4

Ionic bond > Covalent bond > Hydrogen bond > Metallic bond

Relative bond strength



Relative bond to water

Average values of C-C and C-H bond depending on the molecule

C-C bond: 1.54 \AA , 1.54 eV , 100 kJ/mol

Hydrogen bond: 1.15 \AA , 1.15 eV , 10 kJ/mol

Length: 1.54 \AA , 1.15 \AA , 1.15 \AA

Strength: 100 kJ/mol , 10 kJ/mol , 10 kJ/mol

Hydrogen bond: 1.15 \AA , 1.15 eV , 10 kJ/mol

Hydrogen bond: 1.15 \AA , 1.15 eV , 10 kJ/mol

Length: 1.15 \AA , 1.15 \AA , 1.15 \AA

Strength: 10 kJ/mol , 10 kJ/mol , 10 kJ/mol

Hydrogen bond and strong bond

Example



Bond 1 > Bond 2 > Bond 3 > Bond 4



Bond 1 > Bond 2 > Bond 3 > Bond 4



Bond 1 > Bond 2 > Bond 3 > Bond 4



Bond 1 > Bond 2 > Bond 3 > Bond 4

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