

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds

1. Ionic bond
2. Covalent bond
3. Hydrogen bond
4. Metallic bond

There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electron from one atom to another (anions and cations). (Opposite charges).



The difference between ionic and covalent bonding



Covalent bond and polar covalent bond

Covalent bonds are classified as regular (non-polar) or polar.

- Polar bonds are formed by sharing of electrons between atoms.
- Non-polar bonds are formed by sharing of electrons between atoms.
- Very simple bond is a covalent bond.
- Non-polar bonds are formed of two regular and identical atoms.



There are two main categories of covalent bonds

1. Polar covalent bond

Depends on the difference in the electronegativity of atoms. Positive bond is formed by polar and nonpolar.



If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar or non-bonded bond. (Non-polar covalent bond, for the C-C and C-H bonds).

If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond. Nonpolar covalent bond electrons are shared in the C-H and other halogen.

If difference of 1.8 or higher one or both the atoms are more polarized and one bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protophilic Agents



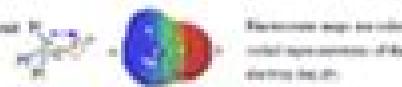
Protophilic and nonphilic agents

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).

1. Nonpolar covalent bond
2. Polar covalent bond
3. Ionic bond



- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector is longer if one has a partial negative charge and the less electronegative atom a partial positive charge (i.e. H_2O dipole, where the oxygen has a partial negative charge).
- In ionic bond, electrons are shared in one atom.
- In ionic bond, electrons are shared in two atoms.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of individual bond dipole moments. If dipole moments are canceled, it just means a neutral dipole.



Sum of vector of the bond dipole for the molecule.

Bond Length and Bond Strength

Relative bond length



Bond: C-C, C-O, C-H, C-F

Length: 1.54 Å, 1.27 Å, 1.47 Å, 1.33 Å

Strength: 100 kJ/mol, 110 kJ/mol, 100 kJ/mol, 110 kJ/mol

Pauli exclusion principle: $\text{C-C} > \text{C-O} > \text{C-H} > \text{C-F}$

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Relative bond strength



Bond: C-C, C-O, C-H, C-F

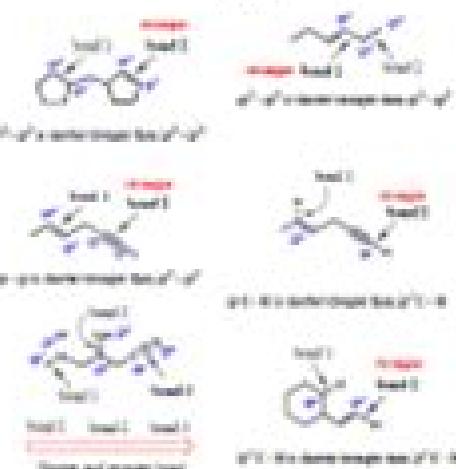
Strength: 100 kJ/mol, 110 kJ/mol, 100 kJ/mol, 110 kJ/mol

Average values of C-C and C-H bond depending on hybridization

| C-C Bond | sp_2-sp_2 | sp_3-sp_3 | $\text{sp}-\text{sp}_3$ |
|---------------|---------------------------|---------------------------|-------------------------|
| Hybridization | sp^2-sp^2 | sp^3-sp^3 | $\text{sp}-\text{sp}^3$ |
| Length | 1.37 Å | 1.46 Å | 1.33 Å |
| Strength | 100 kJ/mol | 110 kJ/mol | 100 kJ/mol |
| Hybridization | sp^2-sp^2 | sp^3-sp^3 | $\text{sp}-\text{sp}^3$ |
| Efficiency | 25 % | 35 % | 50 % |
| Length | 1.39 Å | 1.50 Å | 1.35 Å |
| Strength | 100 kJ/mol | 110 kJ/mol | 100 kJ/mol |

Hybridization reduces bond length

Example



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