

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron). Examples:



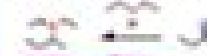
• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bonding and Antibonding Molecular Orbitals



Bonding molecular orbitals



There are two types of molecules

1. Polar molecules



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two multiple bonds consist of one sigma and two pi bonds.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.

• **Polar covalent bonds:**

• **Nonpolar covalent bonds:**

• **Electronegativity scale:**

• If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar covalent bond). Examples: C-C and C-H bonds.

• If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar covalent bond). Examples: C-O and C-N bonds.

• A difference of 1.7 or higher indicates ionic bonding (electron transfer).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



It is a vector quantity.

• The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).

• **δ⁻ (delta minus)** indicates an electron-deficient atom.

• **δ⁺ (delta plus)** indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.



Water is a bent molecule and the dipoles do not cancel.

Bond Length and Bond Strength

• **Bond length:**

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

• **Bond energy:**

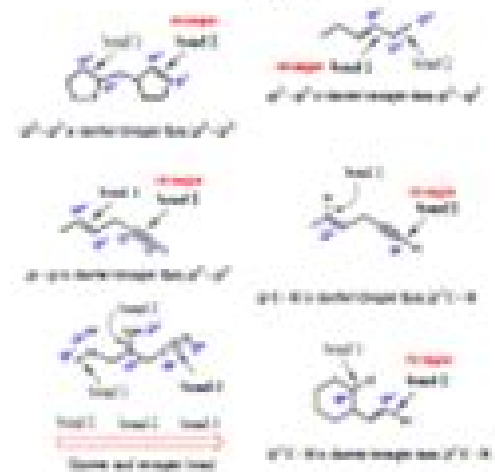


• **Bond energy:**

	sp ³ -C	sp ² -C	sp-C
Hybridization	sp ³	sp ²	sp
Length	154 pm	134 pm	121 pm
Strength	338 kJ/mol	472 kJ/mol	590 kJ/mol
Hybridization	sp ³	sp ²	sp
Length	154 pm	134 pm	121 pm
Strength	338 kJ/mol	472 kJ/mol	590 kJ/mol

• **Bond energy:**

Examples



Organic Chemistry Study Guide

Marc Loudon, Jim Parise



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