

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding between atoms in a molecule.

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (Opposite charges).



Opposites attract and like dissolves like

Opposites attract and like dissolves like (London) bonding of atoms.

- Covalent bonds are formed by sharing of electrons between two atoms.
- Very simple bond is a covalent bond.
- Non-polar covalent bonds consist of two opposites and two like.



Polar bonds are stronger than non-polar.

Polar and Nonpolar Covalent Bonds



- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a covalent bond between chlorine atoms for the C-C and C-Cl bonds).
- If the difference in electronegativity is between 0.5 and 1.7, one atom has a partial negative charge and the other has a partial positive charge (like a polar covalent bond between oxygen and carbon in a carbon dioxide molecule).
- A difference of 1.7 or higher one or both of the atoms have a larger share and a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists

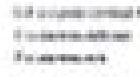


Protagonists and Reactions

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Direction of the dipole moment is from the partial positive to the partial negative.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector is proportional to the partial negative charge and the partial positive charge.
- The dipole plot indicates the charge difference.
- The dipole moment is shown as a distance and angle.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the bond dipole moments. If the dipole moments are equal, it is just simple addition. If not,



Sum of the bond dipole moment for the molecule.

Bond Length and Bond Strength

Reactions of bond length



Length: 0.74 Å, 1.27 Å, 1.42 Å, 1.95 Å.

Strength: 1.00 kcal/mol, 0.75 kcal/mol, 0.75 kcal/mol, 0.75 kcal/mol.

Length and strength are inversely proportional.



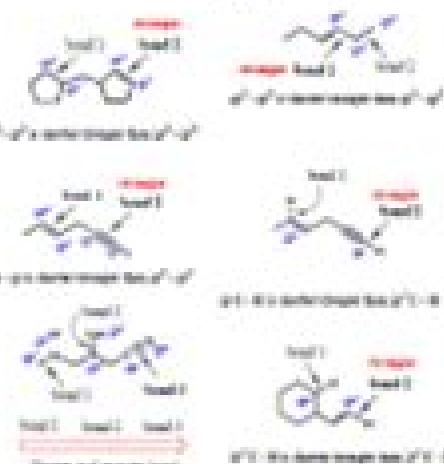
Reactions of bond strength

Average values of C-C and C-H bond depending on hybridization

C-C Bond	sp^2-sp^2	sp^3-sp^3	sp^2-sp^3
Hybridization	sp^2-sp^2	sp^3-sp^3	sp^2-sp^3
Length	1.37 Å	1.54 Å	1.75 Å
Strength	1.00 kcal/mol	0.75 kcal/mol	0.75 kcal/mol
Hybridization	sp^2	sp^3	sp^2-sp^3
Strength	0.75 kcal/mol	0.75 kcal/mol	0.75 kcal/mol
Length	1.37 Å	1.54 Å	1.75 Å
Strength	0.75 kcal/mol	0.75 kcal/mol	0.75 kcal/mol

Higher the bond length and stronger bond

Example



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