

A crystalline solid is formed when 2 or more atoms exchange valence electrons to form an extended stability.

Ionic bond is formed when valence electrons are transferred from one atom to another.

cations = atoms lose electrons to become [+] charged cations.

anions gain electrons to become [-] charged.

In ionic compounds that ions are arranged in a crystal lattice. Strong forces hold the ions together.

Properties of ionic compounds:

- = high melting and boiling points
- = hard & not easily crushed
- = conduct electricity when melted or dissolved because the ions are free to move

covalent bond = electrons are shared. Forming molecules

Covalent compounds have weak forces holding the molecules together.

Properties of covalent compounds:

- low melting and boiling points
- Many covalent compounds are volatile liquids or gases.
- soft, easier to crush
- are not conductors of electricity
- electronegativity = property that tells how strong an

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M Carnoy



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Chek-Chart Farm Tractor Lubrication and Service Guide Chek-Chart Corporation, Chicago,1962 Agricultural Library Notes ,1942 Medicare and Medicaid Guide ,1969 Resources in Education ,1993 **Research in Education** ,1973

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