

A covalent bond = forms when 2 or more atoms exchange valence electrons to increase stability.

ionic bond = forms when valence electrons are transferred from one atom to another

cations = atom loses electrons to become [+] charged cation.

Atom gains electrons to become [-] charged

In ionic compounds that ions are arranged in a crystal lattice. Strong forces hold the ions together.

Properties of ionic compounds:

- = high melting and boiling points
- = hard & not easily crushed
- = conduct electricity when melted or dissolved because the ions are free to move

covalent bond = electrons are shared. Forming molecules

Covalent compounds have weak forces holding the molecules together

Properties of covalent compounds:

- low melting and boiling points
- Many covalent compounds are volatile liquids or gases
- soft, easier to crush
- are not conductors of electricity
- electronegativity = property that tells how strong an

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