

## Honors Chemistry Midterm: Review Packet

## PART 1

1. Write in the number of significant figures for each of the following measurements:

3 3.00 mL    3 0.00290 g    4 50.00 m    2 0.070 kg    1 400 L

2. Using the following masses: 0.0562 kg, 124.213 g and 1635 mg.  $\leftarrow$  convert to grams  
The answer should be reported as: \_\_\_\_\_ g

56.2 g    124.213 g    1.635 g

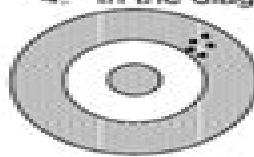
3. a) A 5.75 mL sample of mercury has a measured mass of 77.05 g. The density is 13.49/g

$$D = \frac{m}{v} = \frac{77.05}{5.75 \text{ mL}} =$$

- b) Mercury's accepted density is 13.53 g/mL. The % error in your measurement is: .961%

$$\frac{\text{actual} - \text{exp'd}}{\text{actual}} \times 100 = \frac{13.53 \text{ g/mL} - 13.49 \text{ g/mL}}{13.53 \text{ g/mL}} \times 100$$

4. In the diagram below:



Is this person accurate? no - not in center

Is this person precise? yes - close together

5. Identify the following as a chemical change (CC), physical change (PC), chemical property (CP), or physical property (PP):

a) burning wood CC

b) mass of wood PP

c) flammability of wood CP

d) chopping up wood PC

# Mcps Honors Chemistry B Review Packet Answers

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