

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Covalent Bonds

There are two types of covalent bond: single bond and double bond.

- **Single bond** is formed by the overlap of atomic orbitals (AOs) and is the most common type of covalent bond.



- **Double bond** is formed by the overlap of atomic orbitals (AOs) and is the most common type of covalent bond.



The Relative Strength of Covalent Bonds



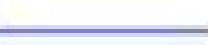
Bond and Bond Energy



Bond Energy



There are two types of covalent bond: single bond and double bond.



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bonds** are formed by the head-on overlap of atomic orbitals (AOs).

• **Pi bonds** are formed by the side-on overlap of atomic orbitals (AOs).

- **Single bond** is a sigma bond.

- **Double bond** is composed of one sigma and one pi bond.



• **Single bond** is stronger than pi bond.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of the atoms, covalent bonds can be polar or nonpolar.



- If the difference in electronegativity is less than 0.5, the bond is nonpolar.

- If the difference in electronegativity is between 0.5 and 1.7, the bond is polar.

- If the difference in electronegativity is greater than 1.7, the bond is ionic.

- A difference of 1.7 or higher can be high for the difference in electronegativity and is a ionic bond.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

- The more electronegative atom has a partial negative charge (δ⁻).

- δ⁻ (delta minus) indicates a partial negative charge.

- δ⁺ (delta plus) indicates a partial positive charge.

- δ⁺ (delta plus) indicates a partial positive charge.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The polarity of the molecule is determined by the vector sum of the dipole moments. A polar molecule has a net dipole moment.



• **Single bond** is stronger than pi bond.

Bond Length and Bond Strength

• **Single bond** is stronger than pi bond.

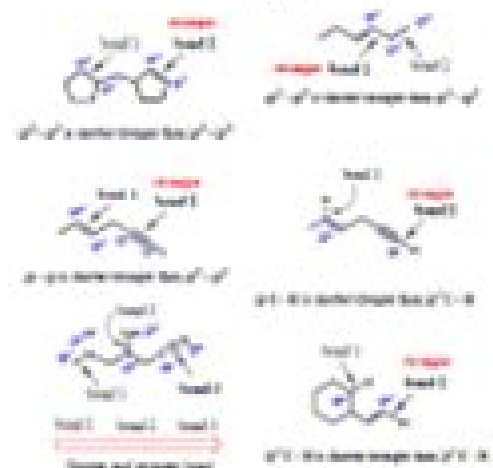


• **Single bond** is stronger than pi bond.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp ³ -sp ³	sp ² -sp ²	sp-sp
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol

• **Single bond** is stronger than pi bond.

Examples



Pavia Organic Chemistry Lab Study Guide

Annelies Wilder-Smith



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