

Name: Marcie HallPeriod: 3Date: 5/1/23PENDA  
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## Note-Taking Guide (CER)

Activity Title: Aquatic Ecosystems- Developing

| Activity Scoreboard<br>when finished, circle highest score achieved |   |   |   |   |   |
|---|---|---|---|---|---|
| Attempt #   | 1 | 2 | 3 | 4 | 5 |
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## Claim:

There are different aquatic systems with diverse organisms.

| Evidence  | Reasoning/Justification   |
|---|---|
| Aquatic ecosystems have abiotic - temp, light, salinity, depth                        | Different organisms can tolerate different factors  |
| Some organisms prefer fresh water vs. saltwater. Bull sharks + manatees tolerate both | Bull sharks and manatees are very different, but tolerate same conditions                 |
| Depth reduces light and increases pressure. Organisms deal with this differently.     | Many types of adaptations exist to survive in these conditions                            |
| The geography of an aquatic environment can determine the life forms present.         | Continental shelf gets lots of light, so there's many living things. Not so in the abyss. |
| Photosynthesis takes place in the neritic zone.                                       | It is the most productive area of the ocean. Plankton are base of food chain.             |
| Metabolism is affected by temp. Many species are cold-blooded.                        | Organisms can migrate or must have adaptations to survive.                                |

# Note Taking Guide For Cells

**Hussin A.Rothana**



## **Note Taking Guide For Cells:**

## Whispering the Strategies of Language: An Psychological Quest through **Note Taking Guide For Cells**

In a digitally-driven world wherever displays reign supreme and quick connection drowns out the subtleties of language, the profound secrets and mental nuances concealed within phrases frequently get unheard. However, situated within the pages of **Note Taking Guide For Cells** a captivating literary treasure pulsing with organic emotions, lies an extraordinary journey waiting to be undertaken. Written by a skilled wordsmith, this wonderful opus attracts readers on an introspective trip, gently unraveling the veiled truths and profound affect resonating within the very cloth of each and every word. Within the psychological depths of this moving evaluation, we will embark upon a heartfelt exploration of the book is primary styles, dissect their captivating writing type, and yield to the effective resonance it evokes heavy within the recesses of readers hearts.

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web answer  $\text{Na}^+ \text{aq}$  and  $\text{SO}_4^{2-} \text{aq}$  identify the chemical equation the spectator ions  $\text{KOH} \text{aq}$   $\text{HNO}_3$   $\text{KNO}_3 \text{aq}$   $\text{H}_2\text{O}$   $\text{L}$   $\text{K}^+ \text{aq}$   $\text{OH}^- \text{aq}$   $\text{H}^+ \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{K}^+ \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{H}_2\text{O}$   $\text{L}$  answer  $\text{K}^+ \text{aq}$  and  $\text{NO}_3^- \text{aq}$  identify the chemical equation the spectator ions  $\text{Pb}^{2+} \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{PbI}_2 \text{s}$   $2\text{KNO}_3 \text{aq}$

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction  $2\text{Ag}^+ \text{aq} + \text{CrO}_4^{2-} \text{aq} \rightarrow \text{Ag}_2\text{CrO}_4 \text{s}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is  $2\text{Na}^+ \text{aq} + 2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + \text{SO}_4^{2-} \text{aq} \rightarrow 2\text{Na}^+ \text{aq} + \text{SO}_4^{2-} \text{aq} + \text{CuCl}_2 \text{s}$  the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web the six  $\text{NO}_3^- \text{aq}$  ions and the six  $\text{Na}^+ \text{aq}$  ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$  total ionic  $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$  net ionic  $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$  the spectator ions in  $\text{KOH}(\text{aq})$  and  $\text{HNO}_3(\text{aq})$  are  $\text{K}^+(\text{aq})$  and  $\text{NO}_3^-(\text{aq})$  how to identify and list the

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web 18 terms lourdes l22 preview terms in this set 4 molecular equation showing the complete chemical formulas of reactions and products  $\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{KCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{KNO}_3(\text{aq})$  complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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web the  $\text{Ag}^+(\text{aq})$  and  $\text{Cl}^-(\text{aq})$  ions become  $\text{AgCl}(\text{s})$  but the  $\text{Na}^+(\text{aq})$  ions and the  $\text{NO}_3^-(\text{aq})$  ions stay as  $\text{Na}^+(\text{aq})$  ions and  $\text{NO}_3^-(\text{aq})$  ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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