

**Using Hess's Law to Determine Enthalpy Change**  
**(Student textbook page 316)**

41. Nitrogen dioxide,  $\text{NO}_2(\text{g})$ , is an emission resulting from the burning of gasoline in an automobile engine that contributes to the formation of smog and acid rain. It can be converted to dinitrogen tetroxide as shown below:



- a. Use Hess's law and the following equations to determine the enthalpy change for this reaction.



- b. Write the thermochemical equation for the overall reaction.

**What Is Required?**

You need to determine the enthalpy change of a reaction by using two other reactions.

**What Is Given?**

You are given the overall equation. You are given two equations and their corresponding enthalpy changes that can be manipulated to represent the individual steps that occur to reach the overall equation.

Plan Your Strategy	
a. Apply Hess's law. The given equations can be manipulated so that their reactants and products match the reactants and products in the overall equation.	
To manipulate the equations, reverse (1).	
When an equation is reversed, the sign of $\Delta H^\circ$ is reversed. If the stoichiometric coefficients in a given equation are multiplied by an integer or a fraction, multiply $\Delta H^\circ$ for that equation by the same integer or fraction.	
Act on Your Strategy	
$-1 \times (1):$	$2\text{NO}_2(\text{g}) \rightarrow \text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \quad \Delta H^\circ = -66.4 \text{ kJ}$
(2):	$\text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{N}_2\text{O}_4(\text{g}) \quad \Delta H^\circ = 11.1 \text{ kJ}$
Plan Your Strategy	
To determine the overall reaction, add the manipulated equations together, cancelling species that are common to both sides.	
Add the $\Delta H^\circ$ values together to find the enthalpy change for the overall reaction.	

# Mhr Chemistry 12 Solutions

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