

NCERT Solutions for Class-XI Chemistry

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Chapter-2

NCERT Chemistry Class 11

1. Calculate the molecular mass of the following:

- Calculate the number of electrons which will together weigh one gram.
- Calculate the mass and charge of one mole of electrons.

(i) Mass of one electron = 9.10939×10^{-31} kg

∴ Number of electrons that weigh 9.10939×10^{-31} kg = 1

∴ Number of electrons that will weigh 1 g (1×10^{-3} kg)

$$= \frac{1}{9.10939 \times 10^{-31} \text{ kg}} \times (1 \times 10^{-3} \text{ kg})$$
$$= 0.1098 \times 10^{31} + 3.1$$
$$= 0.1098 \times 10^{28}$$
$$= 1.098 \times 10^{27}$$

(ii) Mass of one electron = 9.10939×10^{-31} kg

Mass of one mole of electron = $(6.022 \times 10^{23}) \times (9.10939 \times 10^{-31} \text{ kg})$
= 5.48×10^{-7} kg

Charge on one electron = 1.6022×10^{-19} coulomb

Charge on one mole of electron = $(1.6022 \times 10^{-19} \text{ C}) (6.022 \times 10^{23})$
= $9.65 \times 10^4 \text{ C}$

2. (i) Calculate the total number of electrons present in one mole of methane.

(ii) Find (a) the total number and (b) the total mass of neutrons in 7 mg of ^{14}C . (Assume that mass of a neutron = 1.675×10^{-27} kg).

(iii) Find (a) the total mass of protons in 34 mg of NH_3 at STP.

Will the answer change if the temperature and pressure are changed?

2. No the answer will not vary with the change in temperature and pressure because the number of subatomic particles like protons, neutrons and electrons is fixed for each and every element and it does not vary with temperature and pressure.

3. How many neutrons and protons are there in the following nuclei?

^{12}C , ^{16}O , ^{24}Mg , ^{56}Fe , ^{88}Sr

3. ^{13}C : Atomic mass = 13

Atomic number = Number of protons = 6

Number of neutrons = (Atomic mass) - (Atomic number)

$$= 13 - 6 = 7$$

^{16}O :

Atomic mass = 16

Atomic number = 8

Number of protons = 8

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Priti Singhal

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CBSE Class 11 Chemistry Story in English Chapter 2 Priti Singhal, 2024-11-14 This book is structured to align with the latest syllabus and curriculum guidelines, ensuring that the content is both relevant and rigorous. Each chapter begins with a clear set of learning objectives, providing a roadmap for students to understand what they will achieve by the end of the chapter. We have included numerous diagrams, illustrations, and real-life examples to make complex concepts more accessible and engaging.

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