

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Covalent Bonds

There are two types of covalent bond: single bond and double bond.

- **Single bond** is formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end overlap).



- **Double bond** is formed by sharing of electron between two atoms.



The Relative Strength of Covalent Bonds



Single and Double Bonds



Relative Strength



There are two types of covalent bond

Definition of covalent bond



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bond** is formed by head-on overlap of orbitals.

- **pi bond** is formed by side-on overlap of orbitals.

- **Single bond** is a sigma bond.

- **Double bond** consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electron is shared equally (nonpolar bond). For example, the C-C and C-H bonds.

- If the difference in electronegativity is between 0.5 and 1.7, the bond is a polar covalent bond. For example, the C-O and C-N bonds.

- A difference of 1.7 or higher is ionic, and the electron is not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

- The more electronegative atom has a partial negative charge (δ^-), and the less electronegative atom has a partial positive charge (δ^+).

- δ^- (delta minus) indicates an electron-deficient atom.

- δ^+ (delta plus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments.

The geometry of the molecule is important. Symmetrical molecules are nonpolar, even if the dipole moments are present. A good example is carbon dioxide.



Cancel the vector as the other pointing point in the same line.

Bond Length and Bond Strength

The longer the bond, the weaker the bond.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds, and triple bonds are stronger than double bonds.

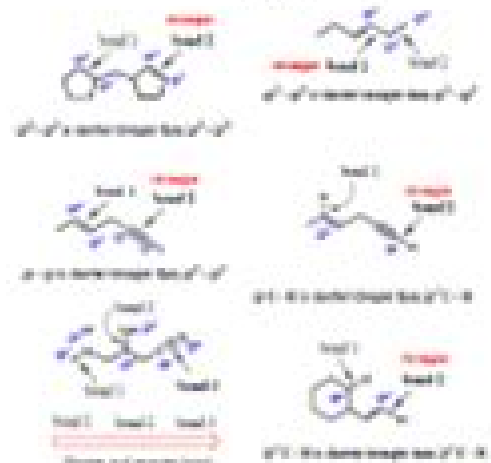


The longer the bond, the weaker the bond.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^2	sp^2-sp	$sp-sp$
C-H Bond	109 pm	108 pm	106 pm
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol

Higher s character, stronger and shorter bond

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 - The Rise of Digital Reading Organic Chemistry With Biological Applications Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry With Biological Applications Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry With Biological Applications Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chemistry With Biological Applications Study Guide
 - Personalized Recommendations
 - Organic Chemistry With Biological Applications Study Guide User Reviews and Ratings

- Organic Chemistry With Biological Applications Study Guide and Bestseller Lists
- 5. Accessing Organic Chemistry With Biological Applications Study Guide Free and Paid eBooks
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 - Participating in Virtual Book Clubs
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 - Creating a Diverse Reading Collection Organic Chemistry With Biological Applications Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Organic Chemistry With Biological Applications Study Guide
 - Setting Reading Goals Organic Chemistry With Biological Applications Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Organic Chemistry With Biological Applications Study Guide
 - Fact-Checking eBook Content of Organic Chemistry With Biological Applications Study Guide
 - Distinguishing Credible Sources

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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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closer to the ridge are younger rift crack at the center of the ridge

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