

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bonding

There are two types of covalent bonding: **sigma bonding** and **pi bonding**.

- Sigma bonds** are formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end) and are **stronger**.



- Pi bonds** are formed by the overlap of atomic orbitals (AOs) in a side-on manner (side-to-side) and are **weaker**.



The Relative Strength of σ and π Bonds



Bond and Bond Energy



Bond Energy



There are two types of covalent bonding

Definition of covalent bond



Sigma and pi bonds

Covalent bonds are classified into **sigma** and **pi** bonds.

- Sigma bonds** are formed by the head-on overlap of atomic orbitals (AOs).
- Pi bonds** are formed by the side-on overlap of atomic orbitals (AOs).
- Sigma bonds** are **stronger** than pi bonds.
- Sigma bonds** are **formed** by the head-on overlap of atomic orbitals (AOs).



σ bonds are stronger than π bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of the atoms, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the bond is **nonpolar**.
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar**.
- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.

- A difference of **0.5 or higher** is **highly polar** and the difference is **very high** if it is **1.7 or higher**.

- A difference of **0.5 or higher** is **highly polar** and the difference is **very high** if it is **1.7 or higher**.

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).



- The dipole moment is a **vector** pointing from the **less electronegative** atom to the **more electronegative** atom.
- The **size** of the dipole moment is a **measure** of the **polarity** of the bond.
- μ = (charge) × (distance)** between the two atoms.
- μ = (charge) × (distance)** between the two atoms.

Molecular Dipole

To determine the **molecular dipole**, the **vector** sum of the **dipole** is used.

The **polarity** of the molecule is **indicated** by the **direction** of the **dipole** moment.



Draw the **dipole** for the **molecule** and **indicate** the **direction**.

Bond Length and Bond Strength

The **length** of the bond is **related** to the **strength** of the bond.



The **length** of the bond is **related** to the **strength** of the bond.

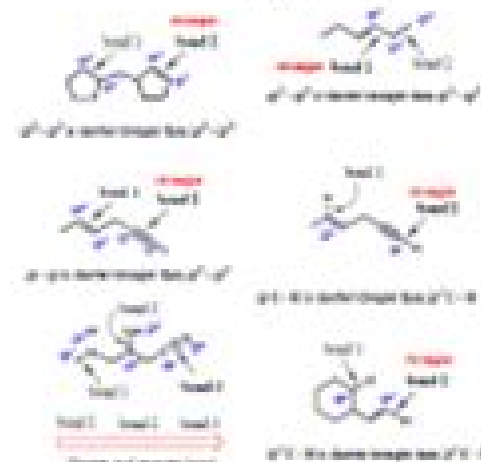


The **length** of the bond is **related** to the **strength** of the bond.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
C-H Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	1.09 Å	1.06 Å	1.03 Å
Strength	99 kcal/mol	112 kcal/mol	120 kcal/mol

Higher s character = stronger and shorter bond

Examples



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Judd E. Hollander



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