

Oxidation Numbers

Summary: An atom connected to a **more** eN atom = +1

An atom connected to a **less** eN atom = -1

An atom connected the **same** atom = 0

* Number of bonds must be taken into account

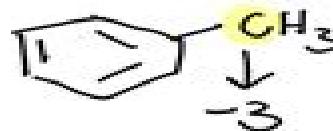
* Charge is also incorporated into oxidation number

Oxidation number \uparrow = Oxidized = reducing agent

Oxidation number \downarrow = reduced = oxidizing agent

Examples: $O=C=O$

+4



Oxidation Reactions

Summary: 1° alcohol $\xrightarrow{[Co]}$ aldehyde $\xrightarrow{[Co] H_2O}$ carboxylic acid



1° alcohol $\xrightarrow{[Co]}$ aldehyde must be dry

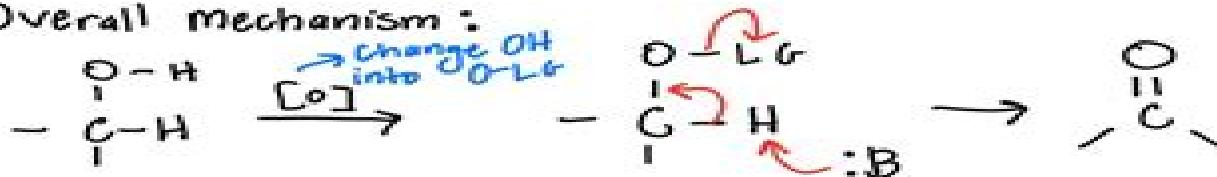
2° alcohol $\xrightarrow{[Co]}$ ketone can be either wet or dry



3° alcohol $\xrightarrow{[Co]}$ no further oxidation



Overall mechanism:



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Harry Linn Fisher

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Table of Contents Organic Chem Phenol Study Guide

1. Understanding the eBook Organic Chem Phenol Study Guide
 - The Rise of Digital Reading Organic Chem Phenol Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chem Phenol Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chem Phenol Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Organic Chem Phenol Study Guide
 - Personalized Recommendations
 - Organic Chem Phenol Study Guide User Reviews and Ratings
 - Organic Chem Phenol Study Guide and Bestseller Lists

5. Accessing Organic Chem Phenol Study Guide Free and Paid eBooks
 - Organic Chem Phenol Study Guide Public Domain eBooks
 - Organic Chem Phenol Study Guide eBook Subscription Services
 - Organic Chem Phenol Study Guide Budget-Friendly Options
6. Navigating Organic Chem Phenol Study Guide eBook Formats
 - ePUB, PDF, MOBI, and More
 - Organic Chem Phenol Study Guide Compatibility with Devices
 - Organic Chem Phenol Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Organic Chem Phenol Study Guide
 - Highlighting and Note-Taking Organic Chem Phenol Study Guide
 - Interactive Elements Organic Chem Phenol Study Guide
8. Staying Engaged with Organic Chem Phenol Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chem Phenol Study Guide
9. Balancing eBooks and Physical Books Organic Chem Phenol Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chem Phenol Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chem Phenol Study Guide
 - Setting Reading Goals Organic Chem Phenol Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Organic Chem Phenol Study Guide
 - Fact-Checking eBook Content of Organic Chem Phenol Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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