

Chapter 7 – Alkenes: Structure and Reactivity

Solutions to Problems

- 7.1 Because two hydrogens must be removed from a saturated compound to introduce an unsaturation, a compound's degree of unsaturation refers to the number of pairs of hydrogens by which its formula differs from that of the corresponding saturated compound. For example, a saturated alkane with four carbons has the formula C_4H_{10} . The compound in (a), C_4H_8 , which has two fewer (or one pair fewer) hydrogens, may have a double bond or a ring. C_4H_{14} thus has a degree of unsaturation of 1.

Compound	Degree of Unsaturation	Structures		
(a) C_4H_8	1	$CH_3CH=CHCH_3$	$CH_3CH_2CH=CH_2$	$(CH_3)_2C=CH_2$
				
(b) C_4H_6	2	$CH_2=CHCH=CH_2$	$CH_3CH=C=CH_2$	$CH_3C\equiv CCH_3$
		$CH_3CH_2C\equiv CH$		
				
(c) C_3H_4	2	$H_2C=C=CH_2$	$CH_3C\equiv CH$	

- 7.2 Unlike the hydrocarbons in the previous problems, the compounds in this problem contain additional elements. Review the rules for these elements.
- Subtract one hydrogen for each nitrogen present to find the formula of the equivalent hydrocarbon – C_6H_4 . Compared to the alkane C_6H_{14} , the compound of formula C_6H_4 has 10 fewer hydrogens, or 5 fewer hydrogen pairs, and has a degree of unsaturation of 5.
 - $C_6H_5NO_2$ also has 5 degrees of unsaturation because oxygen does not affect the equivalent hydrocarbon formula of a compound.
 - A halogen atom is equivalent to a hydrogen atom in calculating the equivalent hydrocarbon formula. For $C_8H_5Cl_3$, the equivalent hydrocarbon formula is C_8H_{12} , and the degree of unsaturation is 3.
 - $C_9H_{16}Br_2$ – one degree of unsaturation.
 - $C_{10}H_{12}N_2O_3$ – 6 degrees of unsaturation.
 - $C_{20}H_{30}ClN$ – 5 degrees of unsaturation.

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 - The Rise of Digital Reading Organic Chemistry 10e Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Organic Chemistry 10e Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Organic Chemistry 10e Solutions Manual
 - User-Friendly Interface

4. Exploring eBook Recommendations from Organic Chemistry 10e Solutions Manual
 - Personalized Recommendations
 - Organic Chemistry 10e Solutions Manual User Reviews and Ratings
 - Organic Chemistry 10e Solutions Manual and Bestseller Lists
5. Accessing Organic Chemistry 10e Solutions Manual Free and Paid eBooks
 - Organic Chemistry 10e Solutions Manual Public Domain eBooks
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 - ePub, PDF, MOBI, and More
 - Organic Chemistry 10e Solutions Manual Compatibility with Devices
 - Organic Chemistry 10e Solutions Manual Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Organic Chemistry 10e Solutions Manual
 - Highlighting and Note-Taking Organic Chemistry 10e Solutions Manual
 - Interactive Elements Organic Chemistry 10e Solutions Manual
8. Staying Engaged with Organic Chemistry 10e Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Organic Chemistry 10e Solutions Manual
9. Balancing eBooks and Physical Books Organic Chemistry 10e Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Organic Chemistry 10e Solutions Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Organic Chemistry 10e Solutions Manual
 - Setting Reading Goals Organic Chemistry 10e Solutions Manual
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Organic Chemistry 10e Solutions Manual
 - Fact-Checking eBook Content of Organic Chemistry 10e Solutions Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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