

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactants.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



$$\text{O}_2 = 32\text{ g} = 32\text{ g}$$

The molar masses of Cl_2 , NO_2 , and NO_3 are 71.00 g/mol, 63.01 g/mol, and 83.01 g/mol, respectively.

Amount of Cl_2	Amount of NO_2	Amount of NO_3	Identifying Reactant	Amount and Name of Limiting Reactant
1 molecule	1 molecule	1 molecule	none	none
4 molecules	4 molecules	4 molecules	NO_2	4 molecules NO_2
7 molecules	8 molecules	6 molecules	Cl_2	6 molecules NO_2 and 7 molecules NO_3
1.00 mol	1.00 mol	1.00 mol	NO_2	1.00 mol NO_2
1.00 mol	0.50 mol	1.00 mol	NO_2	0.50 mol NO_2 and 1.00 mol O_3
1.00 mol	0.00 mol	1.00 mol	NO_2	1.00 mol NO_2 and 1.00 mol O_3
1.00 mol	0.00 mol	1.00 mol	O_3	1.00 mol NO_2 and 1.00 mol O_3
0.000 mol	0.000 mol	1.00 mol	NO_2	1.00 mol NO_2 and 1.00 mol O_3
10.00 g	10.00 g	10.00 g	NO_2	10.00 g NO_2
10.00 g	10.00 g	10.00 g	O_3	10.00 g NO_2 and 10.00 g O_3
10.00 g	10.00 g	20.76 g	O_3	10.00 g NO_2 and O_3

balanced equation mole ratio = 2 mol NO /1 mol O_3

$$10.00 \text{ g} \text{O}_3 \times 1 \text{ mol O}_3/32.00 \text{ g-O}_3 = 0.3125 \text{ mol O}_3$$

$$20.76 \text{ g} \text{NO} \times 1 \text{ mol NO}/46.01 \text{ g-NO} = 0.4464 \text{ mol NO}$$

actual mole ratio = 0.4464 mol NO/0.3125 mol $\text{O}_3 = 2.132$ mol NO/1,000 mol O_3 Because the actual mole ratio of NO : O_3 is larger than the balanced equation mole ratio of NO : O_3 , there is an excess of NO ; O_3 is the limiting reactant.Mass of NO used = 0.3125 mol $\text{O}_3 \times 2$ mol NO/1 mol $\text{O}_3 = 0.6250$ mol NO

$$0.6250 \text{ mol NO} \times 46.01 \text{ g NO}/1 \text{ mol NO} = 28.76 \text{ g NO}$$

Mass of NO_2 produced = 0.6250 mol $\text{NO}_2 \times 46.01 \text{ g NO}_2/1 \text{ mol NO}_2 = 28.76 \text{ g NO}_2$ Excess $\text{NO} = 20.00 \text{ g NO} - 18.76 \text{ g NO} = 1.24 \text{ g NO}$

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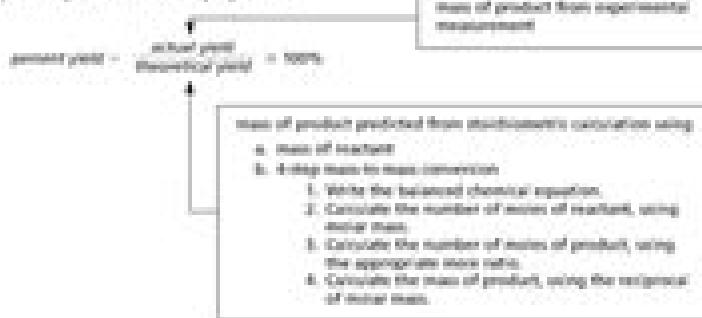
CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a substance used in the manufacturing of some semiconductor devices.

In one experiment, the student yielded 7.47 g of the oxide from a 20.0-g sample of palladium. Determine the percent yield of this reaction. The molar masses of KlO_3 and Na_2PdCl_4 are 108.73 g/mol and 167.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: $7.47 \text{ g K}_2\text{PdCl}_6$
- mass of reactant: 7.00 g KlO_3 correct
- number of moles of reactant: $7.00 \text{ g} \text{KlO}_3 \times \frac{1 \text{ mol KlO}_3}{108.73 \text{ g} \text{KlO}_3} = 0.0644 \text{ mol KlO}_3 = 7.00 \text{ g} \text{KlO}_3 / 1 \text{ mol KlO}_3 = 0.0644 \text{ mol KlO}_3$
- number of moles of product: $0.0644 \text{ mol} \text{KlO}_3 \times \frac{1 \text{ mol K}_2\text{PdCl}_6}{1 \text{ mol KlO}_3} = 0.0644 \text{ mol} \text{K}_2\text{PdCl}_6 = 0.0644 \text{ mol} \text{KlO}_3 / 2 \text{ mol K}_2\text{PdCl}_6 = 0.0322 \text{ mol} \text{K}_2\text{PdCl}_6$
- theoretical yield: $0.0322 \text{ mol} \text{K}_2\text{PdCl}_6 \times \frac{167.44 \text{ g} \text{K}_2\text{PdCl}_6}{1 \text{ mol} \text{K}_2\text{PdCl}_6} = 5.44 \text{ g} \text{K}_2\text{PdCl}_6$ correct
- percent yield: $\frac{7.47 \text{ g} \text{K}_2\text{PdCl}_6}{5.44 \text{ g} \text{K}_2\text{PdCl}_6} \times 100\% = 137\%$

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Barbara Janson Cohen, Memmler

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