

The Rate Equation

CARD
9.1

SPHC
9.1.1 to 9.1.3

The rate of a reaction is measured by observing the rate of change of quantities of reactants or products over time.

The Rate Equation

For a general reaction $A + B \rightarrow C$, where both A and B affect rate:

$$\text{Rate} = k[A]^x[B]^y$$

Where the rate constant, k , is a proportionality constant, which is equal to the rate of the reaction when the concentration of all species in the rate equation is 1 mol/litre, and the exponents, x and y , are the orders of the reaction with respect to A and B respectively.

The Rate Equation

CARD
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9.1.1 to 9.1.3

Experiment	Initial (A) / mol dm ⁻³	Initial (B) / mol dm ⁻³	Initial rate / mol dm ⁻³ s ⁻¹
1	0.1	0.1	0.001
2	0.2	0.1	0.004
3	0.1	0.2	0.002
4	0.2	0.2	0.008

Compare 1 & 2: $[\text{A}]$ doubled in exp 2, $[\text{B}]$ doubled in exp 2. Compare 2 & 3: $[\text{B}]$ doubled in exp 3. Compare 1 & 4: we first take into account the affected rate constant can use the affected rate constant to find the order of B. $k = \text{Rate} / [\text{A}]^1 [\text{B}]^1$
 $0.001 = 10^{-3} \times 1 \cdot 5^1 \times 0.1 = 1.5 \times 10^{-3}$
 $0.008 = 10^{-3} \times 2 \cdot 5^1 \times 0.2 = 2 \cdot 5 \times 10^{-3}$
 $2 \cdot 5 \times 10^{-3} / 1.5 \times 10^{-3} = 2$ (mark) Rate is doubled.

Finding k

SPHC
9.1.2c

Overall reaction is second order with respect to NO_2 and zero order with respect to CO_2 .

Rate equation for this reaction:

k_1^2 (mark)

(1 mark)

Half order of the reaction:

k_2

(1 mark)

Rate of the rate constant, k , for this reaction:

(1 mark)

Structure of Benzene

SPHC
9.1.4b, 9.1.4c

Benzene consists of a planar hexagon of C atoms, each bonded to 3 atoms of carbon atoms in alternating where 3 of its bonds are double bonds to 2 other carbon atoms and to 1 hydrogen atom.

Hydrogen atoms which form other observed double bonds in 1,3,5 positions.



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