

Thermochemistry

Multiple Choice Questions

B 1. When the three types of energy changes are arranged in order of decreasing molar heat values, what is the correct order?

A. chemical, phase, nuclear
B. nuclear, chemical, phase
C. nuclear, phase, chemical
D. phase, chemical, nuclear

B 2. The temperature remains constant as energy is added to a substance. How may the substance be changing?

A. from a gas to a solid
B. from a liquid to a gas
C. from a liquid to a solid
D. in the amount of kinetic energy

A 3. Which metal requires the most energy to raise 1.00 g of it by 1.00°C?

A. aluminum B. copper C. lead D. nickel

B 4. A substance has a molar heat of combustion of -810.4 kJ/mol. When 0.285 mol of the substance is burned in a calorimeter containing 0.60 kg of water, what is the increase of the water temperature?

A. 0.156°C B. 4.48°C C. 7.89°C D. 12.8°C

B 5. The addition of 954 kJ of heat is required to raise the temperature of 225.0 g of a liquid hydrocarbon from 20.5°C to 45.0°C. What is the heat capacity of this hydrocarbon?

A. 0.94 J/g°C B. 1.73 J/g°C C. 1.00 J/g°C D. 942 J/g°C

C 6. If the molar heat of fusion of sodium is 2.63 kJ/mol, how much energy is needed to melt 180.0 g of solid sodium at its melting point?

A. 263 kJ B. 293 kJ C. 29.4 kJ D. 473 kJ

B 7. Which process is exothermic?

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DP Hallahan

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