

Chemistry I (H) - Unit 02 (Pt. II - Nuclear Chemistry) Study Guide

1) Define, compare, and contrast radiation (pg. 122) to radioactivity (pg. 122).

2) Define AND give an example of a radioisotope (pg. 861).

3) Define nuclear reactions.

- Define nuclear equations (pg. 123/869) -
- Example:

4) Define and provide examples of the following forms of radioactive decay:

- alpha decay (pg. 867) -
 - alpha particle (pg. 123/862) -
 - Example Nuclear Equation:
- beta decay (pg. 867) -
 - beta particle (pg. 123/863) -
 - Example Nuclear Equation:
- gamma radiation (pg. 124) -
 - gamma ray (pg. 124/863) -
 - Example Nuclear Equation:
- Study the Table 24.2 from pg. 861 that we drew in our notes about the properties of decay (penetrating power, mass, etc.)

5) Define AND give an example of transmutation (pg. 865).

6) What do we mean by the words emit or emanate? Give an example.

7) Compare a stable nucleus to that of an unstable nucleus.

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