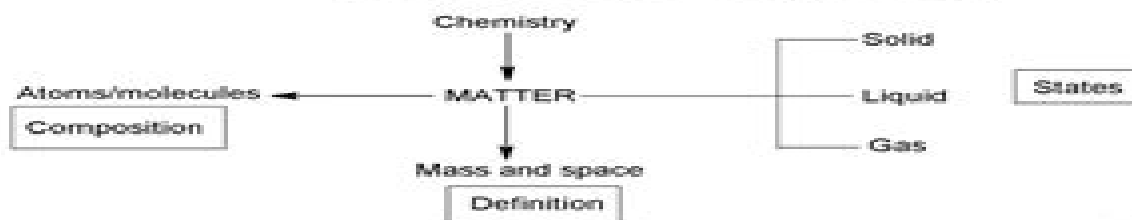


O level Chemistry Notes



Atoms: Building block of matter. It may or may not exist free in nature

Examples: H, He, Na, C

Shells = electrons

Mass = 0

Charge = -1



Nucleus = Protons + Neutrons

Charge = +1 0

Mass = 1 1

Nucleus: It is central part of an atom.

Shells: Circular paths around the nucleus where electrons reside. They are also called orbits or energy levels.

No. of es in a shell = $2n^2$

Where n = no of shell.

1st shell = 2es, 2nd shell = 8es, 3rd shell = 18es.

At least 8 electron should be present in last for the stability of an atom. However if an atom has one shell like Hydrogen and Helium then 2 electrons are required for complete shell and stability.

Atomic number:

Number of protons present in the nucleus of an atom of an element.

Examples: $_{11}\text{Na}$, $_6\text{C}$, $_8\text{O}$

Atomic mass or Mass number or nucleon number: It is the sum of protons and neutrons present in the nucleus of an atom of an element.

Examples: ^{23}Na , ^{12}C , ^{16}O

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Anne T. O'Shea, J. Ron McCartney



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Ground Water Handbook Environmental Protection Agency, U.S., 1992-02 This comprehensive technical guide to assessing and monitoring ground water contamination contains more than 165 charts tables and illustrations and highlights such issues as using models to manage ground water protection programs restoring ground water investigating ground water quality sampling ground water tracing contaminants in the subsurface and monitoring well design and construction

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