

Chemical Reactions: The Nine Bottle Problem

14. The following table summarizes the laboratory experiment. The first question concerns the effect of the concentration of the reactants on the rate of the reaction. The second question concerns the effect of the concentration of the reactants on the equilibrium constant.
- | Run | Initial Concentration of H_2O_2 (M) | Initial Concentration of KI (M) | Initial Concentration of H^+ (M) | Observed Rate (M/s) | Calculated Rate (M/s) |
|-----|---|--|---|---------------------|-----------------------|
| 1 | 0.010 | 0.010 | 0.010 | 0.0001 | 0.0001 |
| 2 | 0.020 | 0.010 | 0.010 | 0.0002 | 0.0002 |
| 3 | 0.010 | 0.020 | 0.010 | 0.0002 | 0.0002 |
| 4 | 0.010 | 0.010 | 0.020 | 0.0001 | 0.0001 |
| 5 | 0.010 | 0.010 | 0.010 | 0.0001 | 0.0001 |
| 6 | 0.010 | 0.010 | 0.010 | 0.0001 | 0.0001 |
- Based on the data, the rate of the reaction is first order with respect to H_2O_2 and first order with respect to KI . The rate of the reaction is zero order with respect to H^+ . The equilibrium constant is 1.0.

Source: <http://www.census.gov/hhes/education/data/tables/2000/2000.ed.education.attainment.attainment.html>

Nine Bottle Problem Answer

Lingsheng Yao



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