

# Reactions of Amines and Epoxides

## Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

## Bond Types, Strength, Length, Polarity and Dipole Moment

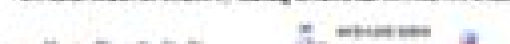
### Free-Radical Reactions

There are two types of free-radical reactions: chain and non-chain reactions.

- Chain reactions are those in which the reaction proceeds by a series of steps, each of which produces a free radical that continues the chain.



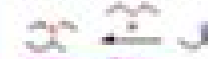
- Non-chain reactions are those in which the reaction proceeds by a single step, and no free radicals are involved.



### The Relative Strength of Covalent Bonds



### Bond and Reaction Rates

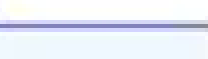


### Bond Energy



### There are two types of covalent bonds: polar and non-polar.

Definition of polar and non-polar bonds:



### Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- Sigma bonds are formed by the end-to-end overlap of atomic orbitals.

• Pi bonds are formed by the side-to-side overlap of atomic orbitals.

- A single bond is a sigma bond.

- A double bond consists of one sigma and one pi bond.



### Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar or non-polar.

- Polar covalent bonds are those in which the electrons are shared unequally between the two atoms, resulting in a dipole moment.

• Non-polar covalent bonds are those in which the electrons are shared equally between the two atoms, resulting in no dipole moment.

• The difference in electronegativity between two atoms is a key factor in determining the polarity of a covalent bond.

• If the difference in electronegativity is between 0.5 and 1.7, the bond is polar covalent.

• If the difference in electronegativity is greater than 1.7, the bond is ionic.

• If the difference in electronegativity is less than 0.5, the bond is non-polar covalent.

• The electronegativity values for the elements are listed in the table below.

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### Dipole Moment

The polarity of a bond is described by the dipole moment ( $\mu$ ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

- The vector dipole moment is a physical quantity that is the product of the charge and the distance between the charges.

- $\mu = q \times r$ , where  $q$  is the charge and  $r$  is the distance between the charges.

- $\mu$  is a vector quantity, and its direction is from the less electronegative atom to the more electronegative atom.

- $\mu$  is a scalar quantity, and its magnitude is the product of the charge and the distance between the charges.

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### Molecular Dipole

To determine the molecular dipole, the vector sum of the bond dipoles is used.

The polarity of the molecule is a physical quantity that is the product of the charge and the distance between the charges.

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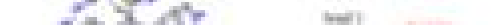
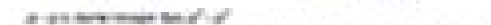
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Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	200 kcal/mol
Hybridization	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
C-H Bond	$sp^3-sp^3$	$sp^2-sp^2$	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	99 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character means stronger bond

### Examples



# Organic Chemistry Laboratory Study Guide

**Ensheng Dong**



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