

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.



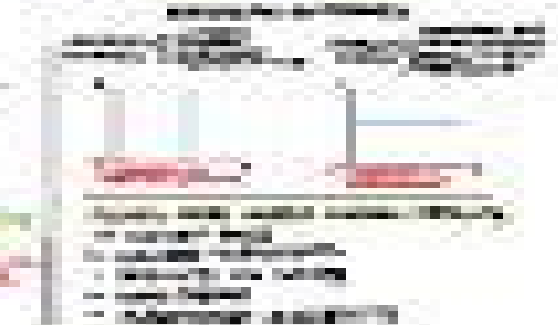
- **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.
 - **Entropy** is a measure of the disorder or randomness of a system.
 - **Heat Engines**: Devices that convert heat into work. The efficiency of a heat engine is given by:

$$\eta = \frac{W}{Q_{in}} = 1 - \frac{Q_{out}}{Q_{in}}$$



- **Third Law of Thermodynamics**: The entropy of a perfect crystal at absolute zero is zero.
 - **Heat Pumps**: Devices that transfer heat from a cold reservoir to a hot reservoir. The coefficient of performance (COP) is given by:

$$COP = \frac{Q_{in}}{W} = \frac{Q_{in}}{Q_{out} - Q_{in}}$$



Microeconomics Final Study Guide

David C. Colander



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