

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 12****STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of  $O_2$ ,  $NO$ , and  $NO_2$  are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of $O_2$	Amount of $NO$	Amount of $NO_2$	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	2 molecules	2 molecules	$NO$	2 molecules $O_2$
2 molecules	4 molecules	4 molecules	$O_2$	2 molecules $NO$
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	2.00 mol	4.00 mol	$NO$	2.00 mol $O_2$
5.00 mol	7.00 mol	7.00 mol	$NO$	1.50 mol $O_2$
1.00 mol	5.00 mol	2.00 mol	$O_2$	4.00 mol $NO$
0.200 mol	0.200 mol	0.200 mol	$NO$	0.400 mol $O_2$
33.00 g	60.02 g	92.02 g	none	none
16.00 g	60.02 g	46.01 g	$O_2$	10.12 g $NO$
16.00 g	30.01 g	28.76 g	$O_2$	1.24 g $NO$

balanced equation mole ratio = 2 mol  $NO$ /1 mol  $O_2$

$16.00 \text{ g } O_2 \times 1 \text{ mol } O_2 / 32.00 \text{ g } O_2 = 0.5125 \text{ mol } O_2$

$20.00 \text{ g } NO \times 1 \text{ mol } NO / 30.01 \text{ g } NO = 0.6664 \text{ mol } NO$

actual mole ratio =  $0.6664 \text{ mol } NO / 0.5125 \text{ mol } O_2 = 1.302 \text{ mol } NO / 1.000 \text{ mol } O_2$

Because the actual mole ratio of  $NO/O_2$  is larger than the balanced equation mole ratio of  $NO/O_2$ , there is an excess of  $NO$ ;  $O_2$  is the limiting reactant.

Mass of  $NO$  used =  $0.5125 \text{ mol } O_2 \times 2 \text{ mol } NO / 1 \text{ mol } O_2 = 0.4250 \text{ mol } NO$

$0.4250 \text{ mol } NO \times 30.01 \text{ g } NO / 1 \text{ mol } NO = 12.76 \text{ g } NO$

Mass of  $NO_2$  produced =  $0.4250 \text{ mol } NO_2 \times 46.01 \text{ g } NO_2 / 1 \text{ mol } NO_2 = 19.56 \text{ g } NO_2$

Excess  $NO = 20.00 \text{ g } NO - 12.76 \text{ g } NO = 7.24 \text{ g } NO$

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**CHAPTER 13****STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yields of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium chloride. Determine the percent yield of this reaction. The molar masses of  $Ga$  and  $Ga_2O_3$  are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield unknown 7.42 g  $Ga_2O_3$
- mass of reactant: 7.00 g  $Ga$  correct
- number of moles of reactant:  $7.00 \text{ g } Ga \times \frac{1 \text{ mol } Ga}{69.72 \text{ g } Ga} = 0.100 \text{ mol } Ga$
- number of moles of product:  $0.100 \text{ mol } Ga \times \frac{2 \text{ mol } Ga_2O_3}{1 \text{ mol } Ga} = 0.200 \text{ mol } Ga_2O_3$
- theoretical yield:  $0.200 \text{ mol } Ga_2O_3 \times \frac{157.44 \text{ g } Ga_2O_3}{1 \text{ mol } Ga_2O_3} = 31.49 \text{ g } Ga_2O_3$  correct
- percent yield:  $\frac{7.42 \text{ g } Ga_2O_3}{31.49 \text{ g } Ga_2O_3} \times 100 = 23.6\%$

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## **Table of Contents Mcgraw Hill Study Guide Answer Science**

1. Understanding the eBook Mcgraw Hill Study Guide Answer Science
  - The Rise of Digital Reading Mcgraw Hill Study Guide Answer Science
  - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Study Guide Answer Science
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mcgraw Hill Study Guide Answer Science
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Study Guide Answer Science
  - Personalized Recommendations
  - Mcgraw Hill Study Guide Answer Science User Reviews and Ratings
  - Mcgraw Hill Study Guide Answer Science and Bestseller Lists
5. Accessing Mcgraw Hill Study Guide Answer Science Free and Paid eBooks
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  - Mcgraw Hill Study Guide Answer Science eBook Subscription Services
  - Mcgraw Hill Study Guide Answer Science Budget-Friendly Options

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  - ePub, PDF, MOBI, and More
  - Mcgraw Hill Study Guide Answer Science Compatibility with Devices
  - Mcgraw Hill Study Guide Answer Science Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mcgraw Hill Study Guide Answer Science
  - Highlighting and Note-Taking Mcgraw Hill Study Guide Answer Science
  - Interactive Elements Mcgraw Hill Study Guide Answer Science
8. Staying Engaged with Mcgraw Hill Study Guide Answer Science
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mcgraw Hill Study Guide Answer Science
9. Balancing eBooks and Physical Books Mcgraw Hill Study Guide Answer Science
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mcgraw Hill Study Guide Answer Science
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Mcgraw Hill Study Guide Answer Science
  - Setting Reading Goals Mcgraw Hill Study Guide Answer Science
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mcgraw Hill Study Guide Answer Science
  - Fact-Checking eBook Content of Mcgraw Hill Study Guide Answer Science
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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