

1. Complete and classify the following partial equations.



2. Given:  $2\text{NaNO}_3 \rightarrow 2\text{Na} + \text{O}_2$  calculate the mass of oxygen gas produced if 145 g of  $\text{NaNO}_3$  is decomposed.

65.4%

3. Calculate the mass of oxygen gas required to burn 240 g of wood ( $\text{C}_5\text{H}_8\text{O}_2$ ).

25.4%

4. If 14.0 moles of butane ( $\text{C}_4\text{H}_{10}$ ) is burned:

a) calculate the number of moles of oxygen gas required. (29.0 mol O<sub>2</sub>)

b) calculate the number of moles of carbon dioxide produced. (14.0 mol CO<sub>2</sub>)

c) calculate the mass of water produced. (10.8 g H<sub>2</sub>O)

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.

a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. (2.7 g Al)

b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and zinc nitrate. (18.2 g Zn)

6. Given:  $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$

a) How many grams of Cu are required to react with 3.00 g of silver nitrate? (0.56 g Cu)

b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient zinc nitrate? (1.50 g Ag)

# Nelson Chemistry 30 Answer Key

**Arthur James Wells**

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