

KEY

1. Complete and classify the following partial equations.

- a) $1 \text{ Fe} + 3 \text{H}_2\text{SO}_4 \rightarrow 1 \text{ FeSO}_4 + 3 \text{H}_2$ (single repl)
 b) $\text{NaI} + \text{Cl}_2 \rightarrow$
 c) $1 \text{ Zn} + 1 \text{ Pb(NO}_3)_2 \rightarrow 1 \text{ Zn(NO}_3)_2 + 1 \text{ Pb}$ (single repl)
 d) $\text{Al}_2\text{O}_3 + \text{HCl} \rightarrow$
 e) $2 \text{ Cu} + \text{O}_2 \rightarrow 2 \text{ CuO}$ or $4 \text{ Cu} + 1 \text{ O}_2 \rightarrow 2 \text{ Cu}_2\text{O}$ (synthesis)
 f) $\text{Al}_2\text{O}_3 \rightarrow$
 g) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ (synth)
 h) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
 i) $1 \text{ Mg} + 2 \text{ HNO}_3 \rightarrow 1 \text{ Mg(NO}_3)_2 + 1 \text{ H}_2$ (s.r.)
 j) $\text{Na} + \text{H}_2\text{O} \rightarrow$
 k) $3 \text{ NaOH} + \text{H}_3\text{PO}_4 \rightarrow 1 \text{ Na}_3\text{PO}_4 + 3 \text{ H}_2\text{O}$ (acid base)
 l) $\text{H}_2\text{SO}_4 + \text{Al(OH)}_3 \rightarrow$
 m) $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$ (combustion)
 n) $\text{AlCl}_3 \rightarrow$
 o) $2 \text{ KOH} + \text{H}_2\text{SO}_4 \rightarrow 1 \text{ K}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$ (acid base)
 p) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
 q) $1 \text{ NaOH} + \text{HCl} \rightarrow 1 \text{ NaCl} + 1 \text{ H}_2\text{O}$ (acid base)
 r) $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
 s) $1 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$ (combustion)
 t) $\text{H}_2\text{O} \rightarrow$
 u) $2 \text{ HNO}_3 + \text{CaCO}_3 \rightarrow 1 \text{ Ca(NO}_3)_2 + 1 \text{ CO}_2 + 1 \text{ H}_2\text{O}$ (s.r.)
 v) $\text{CH}_4 + \text{O}_2 \rightarrow$
 w) $1 \text{ Pb} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ PbSO}_4 + 1 \text{ H}_2$ (s.r.)
 x) $\text{KI} + \text{Br}_2 \rightarrow$
 y) $1 \text{ Ba(OH)}_2 + 2 \text{ HNO}_3 \rightarrow 1 \text{ Ba(NO}_3)_2 + 2 \text{ H}_2\text{O}$ (acid base)
 z) $\text{Pb} + \text{Cu(NO}_3)_2 \rightarrow$

2. Given: $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$. Calculate the mass of oxygen gas produced if 143 g of NaClO_3 is decomposed.

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($\text{C}_{21}\text{H}_{44}$).

254g

4. If 14.0 moles of butane (C_4H_{10}) is burned:

- a) calculate the number of moles of oxygen gas required. 90.0 mol
 b) calculate the number of moles of carbon dioxide produced. 56.0 mol
 c) calculate the mass of water produced. 1080 g

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.

- a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. 27.4g Al
 b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. 182 g Zn

6. Given: $\text{Cu} + 2 \text{ AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{ Ag}$

- a) How many grams of Cu are required to react with 3.00 g of silver nitrate? 0.544g Cu
 b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? 1.70g Ag

2) NaClO_3 1.5675 mol
 $2 \text{ NaClO}_3 \rightarrow 3 \text{ O}_2$
 $1.5675 \text{ mol} \times \frac{3}{2} = 2.35125 \text{ mol O}_2$
 $2.35125 \text{ mol} \times 32 \text{ g/mol} = 75.24 \text{ g}$



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Arthur James Wells



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