

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Enthalpy** (H) is a thermodynamic property that combines internal energy (U) and the product of pressure (P) and volume (V):

$$H = U + PV$$

 - **Enthalpy change** (ΔH) for a process at constant pressure is equal to the heat added to the system:

$$\Delta H = Q_p$$

 - **Enthalpy of formation** (ΔH_f°) is the enthalpy change when one mole of a compound is formed from its elements in their standard states.



- **Free Energy** (G) is a thermodynamic potential that combines enthalpy (H) and entropy (S):

$$G = H - TS$$

 - **Free energy change** (ΔG) determines the spontaneity of a process at constant temperature and pressure:

$$\Delta G < 0$$
 (spontaneous)

$$\Delta G = 0$$
 (equilibrium)

$$\Delta G > 0$$
 (non-spontaneous)

- **Equilibrium** is a state where the forward and reverse rates of a process are equal, and there is no net change in the system.



- **Equilibrium constant** (K) is a measure of the extent of a chemical reaction at equilibrium.



- **Electrochemistry** is the study of the relationship between electrical energy and chemical reactions.

- **Standard cell potential** (E°) is the potential difference between two half-cells under standard conditions.

- **Nernst equation** relates the cell potential to the concentrations of the reactants and products:

$$E = E^\circ - \frac{RT}{nF} \ln Q$$

where E is the cell potential, E° is the standard cell potential, R is the gas constant, T is the temperature, n is the number of moles of electrons transferred, F is the Faraday constant, and Q is the reaction quotient.



- **Electrolysis** is the process of using electrical energy to drive a non-spontaneous chemical reaction.

- **Faraday's laws of electrolysis** describe the relationship between the amount of substance produced and the amount of electricity used.

- **Electrochemical cells** are devices that convert chemical energy into electrical energy.

- **Galvanic cells** are spontaneous electrochemical cells.

- **Electrolytic cells** are non-spontaneous electrochemical cells.



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