

# ATOMIC STRUCTURE

## RUTHERFORD'S MODEL

- Discovered that an atom is mostly empty space with tiny particles.
- Charged particles move in a small elliptical path around a central nucleus.
- This can only hold the negatively charged particles in the center.

## BOHR'S MODEL

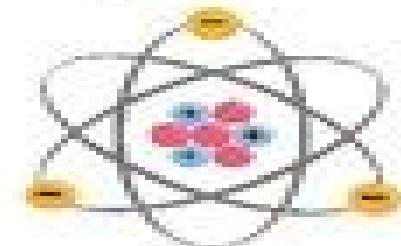
- Shows the elliptical path of an electron.
- Electrons
- Electrons

## THE JELLY PUDDING MODEL



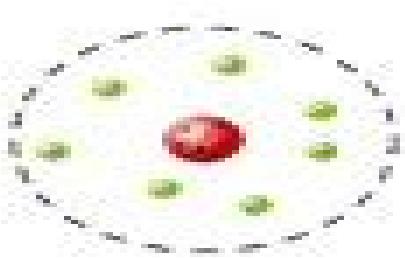
## WAVE-MECHANICAL MODEL

- Shows electrons moving in paths, like tiny clouds like in clouds. Orbitals are regions of space likely where an electron.



## WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
  - PROTONS - positively charged particles
  - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals
  - Electrons DO move around the nucleus in different patterns.



Subatomic Particle	Location in the Atom	Charge	Weight
proton	center (in the nucleus)	positive	1.67 x 10 <sup>-24</sup>
neutron	center (in the nucleus)	neutral (no charge)	1.67 x 10 <sup>-24</sup>
electron	around the nucleus in elliptical orbits	negative	1.67 x 10 <sup>-24</sup> x 10 <sup>-30</sup>

## IONS

- If an atom gains or loses an electron, it becomes an ion.
- This gives the atom a charge:
  - Loss of electrons, the atom becomes positive.
  - Gains of electrons, the atom becomes negative.

## ISOTOPES

- An isotope is when an atom has a different mass number.
- An isotope is still the same element, it is just a little different mass, every atom has all other atoms different in slightly different mass.

# Pearson Chemistry Atomic Structure Study Guide

## Answers

**Bruce Reinhart**

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1. Understanding the eBook Pearson Chemistry Atomic Structure Study Guide Answers
  - The Rise of Digital Reading Pearson Chemistry Atomic Structure Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Atomic Structure Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chemistry Atomic Structure Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Atomic Structure Study Guide Answers

- Personalized Recommendations
- Pearson Chemistry Atomic Structure Study Guide Answers User Reviews and Ratings
- Pearson Chemistry Atomic Structure Study Guide Answers and Bestseller Lists

5. Accessing Pearson Chemistry Atomic Structure Study Guide Answers Free and Paid eBooks

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- Pearson Chemistry Atomic Structure Study Guide Answers eBook Subscription Services
- Pearson Chemistry Atomic Structure Study Guide Answers Budget-Friendly Options

6. Navigating Pearson Chemistry Atomic Structure Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
- Pearson Chemistry Atomic Structure Study Guide Answers Compatibility with Devices
- Pearson Chemistry Atomic Structure Study Guide Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Pearson Chemistry Atomic Structure Study Guide Answers
- Highlighting and Note-Taking Pearson Chemistry Atomic Structure Study Guide Answers
- Interactive Elements Pearson Chemistry Atomic Structure Study Guide Answers

8. Staying Engaged with Pearson Chemistry Atomic Structure Study Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Pearson Chemistry Atomic Structure Study Guide Answers

9. Balancing eBooks and Physical Books Pearson Chemistry Atomic Structure Study Guide Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Pearson Chemistry Atomic Structure Study Guide Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Pearson Chemistry Atomic Structure Study Guide Answers

- Setting Reading Goals Pearson Chemistry Atomic Structure Study Guide Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Pearson Chemistry Atomic Structure Study Guide Answers

- Fact-Checking eBook Content of Pearson Chemistry Atomic Structure Study Guide Answers
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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