

ATOMIC STRUCTURE

RUTHERFORD'S MODEL

- Discovered that an atom is mostly made up of empty space
- Center of the atom is a small dense positively charged nucleus
- This nucleus holds the negatively charged particles in the atom

BOHR'S MODEL

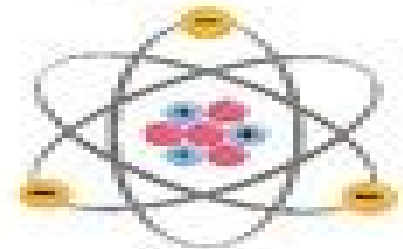
- There are 2 main pieces of an atom:
- Nucleus
- Electrons

THE OLD MODEL OF THE ATOM



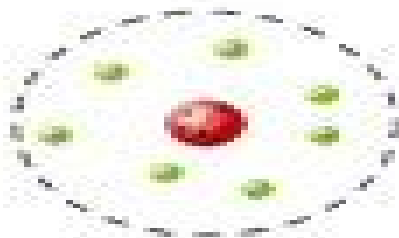
WAVE-MECHANICAL MODEL

- Electrons are moving so fast that they look like a cloud. Orbitals are regions of most likely electron location



WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
 - PROTONS - positively charged particles
 - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals
 - Electrons DO move around the nucleus in different patterns



Subatomic Particle	Location in the Atom	Charge	Weight
proton	center (in the nucleus)	positive	1 amu
neutron	center (in the nucleus)	neutral (no charge)	1 amu
electron	around the nucleus in clouds or orbitals	negative	Almost equal to 1/1836

IONS

- If an atom gains or loses an electron, it is called an ion
- This gives the atom a charge:
 - Lose an electron, the atom becomes positive
 - Gain an electron, the atom becomes negative

ISOTOPES

- An isotope is when an atom is physically different
- An isotope is still the same element, it is just a little different from every other atom of the same element
 - slightly different mass

Pearson Chemistry Atomic Structure Study Guide Answers

Bruce Reinhart



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1. Understanding the eBook Pearson Chemistry Atomic Structure Study Guide Answers
 - The Rise of Digital Reading Pearson Chemistry Atomic Structure Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Atomic Structure Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Atomic Structure Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Atomic Structure Study Guide Answers

- Personalized Recommendations
 - Pearson Chemistry Atomic Structure Study Guide Answers User Reviews and Ratings
 - Pearson Chemistry Atomic Structure Study Guide Answers and Bestseller Lists
5. Accessing Pearson Chemistry Atomic Structure Study Guide Answers Free and Paid eBooks
 - Pearson Chemistry Atomic Structure Study Guide Answers Public Domain eBooks
 - Pearson Chemistry Atomic Structure Study Guide Answers eBook Subscription Services
 - Pearson Chemistry Atomic Structure Study Guide Answers Budget-Friendly Options
 6. Navigating Pearson Chemistry Atomic Structure Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chemistry Atomic Structure Study Guide Answers Compatibility with Devices
 - Pearson Chemistry Atomic Structure Study Guide Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry Atomic Structure Study Guide Answers
 - Highlighting and Note-Taking Pearson Chemistry Atomic Structure Study Guide Answers
 - Interactive Elements Pearson Chemistry Atomic Structure Study Guide Answers
 8. Staying Engaged with Pearson Chemistry Atomic Structure Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Atomic Structure Study Guide Answers
 9. Balancing eBooks and Physical Books Pearson Chemistry Atomic Structure Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Atomic Structure Study Guide Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pearson Chemistry Atomic Structure Study Guide Answers
 - Setting Reading Goals Pearson Chemistry Atomic Structure Study Guide Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pearson Chemistry Atomic Structure Study Guide Answers

- Fact-Checking eBook Content of Pearson Chemistry Atomic Structure Study Guide Answers
- Distinguishing Credible Sources

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- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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