

## 6-Month CDs



# Cd Rates Guide

**Dave Ogershok**

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## **Cd Rates Guide:**

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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction time the total pressure  $p_a = x_a p_{\text{total}}$  mole fraction  $x_a = \frac{\text{moles } a}{\text{total}}$  sum  $n_1 + \dots + x_{n-1}$  mole fraction of substance  $a$  is often denoted by  $x_a$  and

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 30 g benzene s molar mass  $C_6H_6$  128 g mol<sup>-1</sup> 78 g mol<sup>-1</sup> hence  $C_6H_6$  s number of moles  $\frac{30}{78}$  mol 0.3846 mol molar mass of carbon tet  $CCl_4$  154 g mol<sup>-1</sup>

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web jan 30 2020 ans the molality of solution is 1 mol kg<sup>-1</sup> or 1 m example 0.211 g of urea  $NH_2CONH_2$  was dissolved in 100 g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11.11 g mass of solvent water 100 g 0.1 kg to find molarity of solution solution

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web solution moles of  $CH_3OH$  5.532 0.17 mole moles of  $H_2O$  40.182 2.2 moles therefore according to the equation mole fraction of  $CH_3OH$  0.172 0.17 mole fraction of  $CH_3OH$  0.073 to solve more examples on mole fraction formula and practice more question please visit byju s com

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0.449 molal kbr that contains 2.92 kg of water 1 molality moles solute divided by kilograms solute 0.449 mol kg<sup>-1</sup> x 2.92 kg x 1.31108 mol of kbr 2 moles times molar mass equals grams 1.31108 mol 119.0023 g mol 156 g kbr

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web calculate the mass percent mole fraction molality and molarity of hexane  $C_6H_{14}$  d 0.660 g cm<sup>-3</sup> when it is mixed with heptane  $C_7H_{16}$  d 0.684 g cm<sup>-3</sup> in a 35.0550 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of

caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web since the mole fraction of the  $O_2$  is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $O_2$  are mixed in a flask at room temperature

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