

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



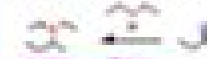
• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding



Bonding energy



There are two types of bonding

Definition of ionic and covalent bonds



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

• **Sigma bonds** are formed by head-on overlap of orbitals.

• **Pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two double bonds consist of one sigma and one pi bond.**



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity of the atoms, covalent bonds can be polar and nonpolar.

• **Polar covalent bonds** are formed by unequal sharing of electrons between two atoms.

• **Nonpolar covalent bonds** are formed by equal sharing of electrons between two atoms.

• If the difference in electronegativity is between 0.5 and 1.7, the bond is polar covalent. Examples are the C-Cl and C-Br bonds.

• If the difference in electronegativity is between 1.7 and 2.0, the bond is ionic. Examples are the Na-Cl and K-Br bonds.

• A difference of 2.0 or higher indicates that the electrons are not shared and an ionic bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



Electronegativity is a measure of the ability of an atom to attract electrons.

• The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).

• **δ⁻ (delta minus)** indicates an electron-deficient atom.

• **δ⁺ (delta plus)** indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.



Water is a bent molecule and the dipoles do not cancel.

Bond Length and Bond Strength

The length of a bond is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds and are shorter.

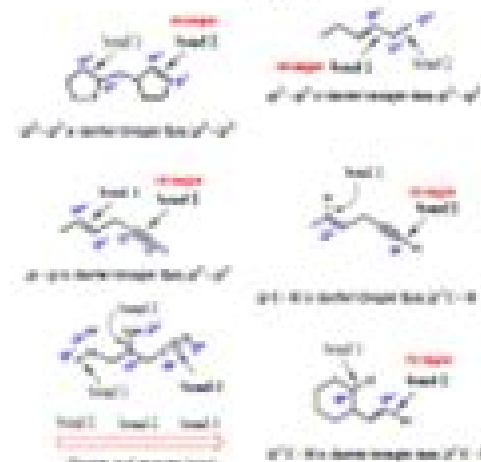


The length of a bond is the distance between the nuclei of the two atoms.

	H ₂ C=CH ₂	H ₂ C≡CH ₂	H ₂ C≡CH
Hybridization	sp ²	sp	sp
Length	134 pm	120 pm	106 pm
Strength	611 kJ/mol	837 kJ/mol	974 kJ/mol
Hybridization	sp ²	sp	sp
Hybridization	sp ²	sp	sp
Length	134 pm	120 pm	106 pm
Strength	611 kJ/mol	837 kJ/mol	974 kJ/mol

Triple bonds are stronger and shorter than double bonds.

Examples



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VM Jensen

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