

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonds

There are two types of bonds: ionic and covalent bonds.

- Ionic bonds are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron gain enthalpy).



- Covalent bonds are formed by sharing of electrons between two atoms.



The Relative Strength of Bonds



Bond and Bond Energy



Bond Energy



There are two types of bonds

Definition of Polarity



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

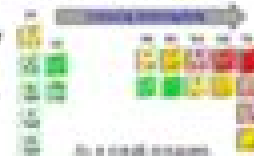
- Sigma bonds are formed by head-on overlap of orbitals.
- pi bonds are formed by side-on overlap of orbitals.
- A single bond is a sigma bond.
- A double bond consists of one sigma and one pi bond.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

- A difference of 1.7 or higher can be high for the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Linear the vector sum of the dipoles is zero.

Bond Length and Bond Strength

The length of the bond is important.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

Double bonds are stronger than single bonds.

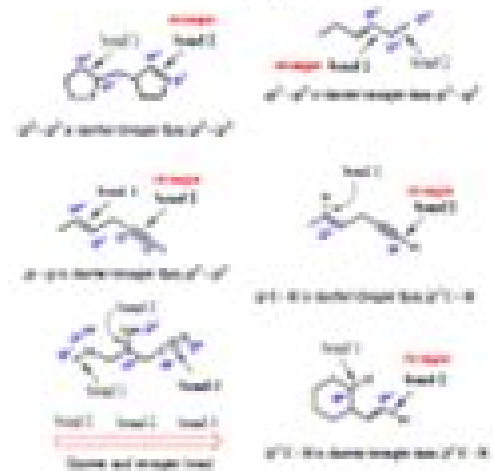


The length of the bond is important.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3 - sp^3	sp^2 - sp^2	sp - sp
C-H Bond	sp^3 - sp^3	sp^2 - sp^2	sp - sp
Length	109 pm	106 pm	103 pm
Strength	413 kJ/mol	439 kJ/mol	505 kJ/mol

Higher s character means shorter and stronger bond

Examples



Pavia Organic Chemistry Lab Study Guide

Ed Bowker Staff, Staff Bowker, Ed



Pavia Organic Chemistry Lab Study Guide:

Organic Laboratory Experiments of the Future Bruce H. Lipshutz, Hubert Muchalski, 2025-12-05 This lab manual offers students the opportunity to learn organic chemistry through a green chemistry lens resulting in a more sustainable future looking course Lipshutz and Muchalski's Organic Laboratory Experiments of the Future Doing Chemistry in Water approaches the subject with experiments diagrams and illustrations shaped by student feedback and honed through years of research and laboratory experience Organic Laboratory Experiments of the Future Doing Chemistry in Water includes material on reactions run in water including but not limited to ppm Pd catalyzed couplings SNAr reactions nitro group reductions enzyme catalyzed reactions catalytic hydrogenation of olefins olefin metathesis reductive aminations and chemoenzymatic sequences This lab manual is ideal for courses in organic chemistry biochemistry sustainable and green chemistry and environmentally responsible lab courses Provides practical information and techniques for utilizing green chemistry and chemistry in water in particular in the study of modern organic chemistry Includes numerous figures examples illustrative problems and appendices that reinforce laboratory concepts and methods Features coverage on experiments that focus on chemocatalysis as well as biocatalysis reactions that include Nobel Prize winning organometallic chemistry multi step sequences involving chemocatalysis or chemoenzymatic catalysis

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Table of Contents Pavia Organic Chemistry Lab Study Guide

1. Understanding the eBook Pavia Organic Chemistry Lab Study Guide
 - The Rise of Digital Reading Pavia Organic Chemistry Lab Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Pavia Organic Chemistry Lab Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pavia Organic Chemistry Lab Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pavia Organic Chemistry Lab Study Guide
 - Personalized Recommendations
 - Pavia Organic Chemistry Lab Study Guide User Reviews and Ratings
 - Pavia Organic Chemistry Lab Study Guide and Bestseller Lists
5. Accessing Pavia Organic Chemistry Lab Study Guide Free and Paid eBooks

- Pavia Organic Chemistry Lab Study Guide Public Domain eBooks
- Pavia Organic Chemistry Lab Study Guide eBook Subscription Services
- Pavia Organic Chemistry Lab Study Guide Budget-Friendly Options
- 6. Navigating Pavia Organic Chemistry Lab Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Pavia Organic Chemistry Lab Study Guide Compatibility with Devices
 - Pavia Organic Chemistry Lab Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pavia Organic Chemistry Lab Study Guide
 - Highlighting and Note-Taking Pavia Organic Chemistry Lab Study Guide
 - Interactive Elements Pavia Organic Chemistry Lab Study Guide
- 8. Staying Engaged with Pavia Organic Chemistry Lab Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pavia Organic Chemistry Lab Study Guide
- 9. Balancing eBooks and Physical Books Pavia Organic Chemistry Lab Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pavia Organic Chemistry Lab Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pavia Organic Chemistry Lab Study Guide
 - Setting Reading Goals Pavia Organic Chemistry Lab Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pavia Organic Chemistry Lab Study Guide
 - Fact-Checking eBook Content of Pavia Organic Chemistry Lab Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web media stutenkerl the raw raisin studded pastry before it is baked and then after baking a stutenkerl belongs to the saint nicholas tradition in the german speaking countries it is a pastry made of stuten sweet leavened dough in the form of a man

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