

Reactions of Amines and Epoxides

Preparation of Amines

Amines are often made by reduction of nitriles or by reductive amination of aldehydes and ketones.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bond

There are two types of covalent bond: **sigma bond** and **pi bond**.

- Sigma bonds** are formed by the overlap of atomic orbitals (AOs) in a head-on manner (end-to-end overlap).



- Pi bonds** are formed by the overlap of atomic orbitals (AOs) in a side-on manner (side-to-side overlap).



The Relative Strength of σ and π Bonds



Bond and Bond Length



Bond Energy



There are two types of covalent bond



Sigma and pi bonds

Covalent bonds are classified into **sigma bonds** and **pi bonds**.

- Sigma bonds** are formed by the head-on overlap of atomic orbitals (AOs).
- Pi bonds** are formed by the side-on overlap of atomic orbitals (AOs).
- Sigma bonds** are stronger than pi bonds.
- Sigma bonds** are formed by the overlap of s, sp, sp², and sp³ hybrid orbitals.



σ bonds are stronger than π bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be **polar** or **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the bond is **nonpolar**.
- If the difference in electronegativity is **between 0.5 and 1.7**, the bond is **polar**.
- If the difference in electronegativity is **greater than 1.7**, the bond is **ionic**.

A difference of 1.7 or higher indicates that the electrons are not shared equally and an ionic bond is formed.

A difference of 0.5 or higher indicates that the electrons are not shared equally and a polar bond is formed.

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).



- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector dipole moment is a product of the charge (q) and the distance (r) between the charges ($\mu = qr$).
- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.
- The dipole moment is a vector pointing from the less electronegative atom to the more electronegative atom.

Molecular Dipole

To determine the molecular dipole, the vector sum of the bond dipoles is used.

The polarity of the molecule is a product of the charge (q) and the distance (r) between the charges ($\mu = qr$).



Sum the vector of the bond dipoles to find the molecular dipole.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.



Double bonds are stronger than single bonds.

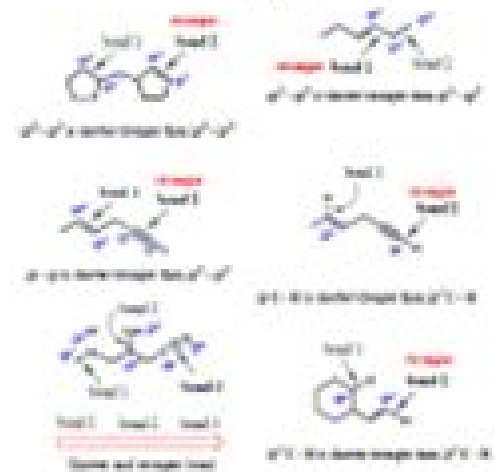


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp ³ -sp ³	sp ² -sp ²	sp ³ -sp ²
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp ³ -sp ²
Length	1.54 Å	1.34 Å	1.47 Å
Strength	347 kJ/mol	461 kJ/mol	377 kJ/mol
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp ³ -sp ²
C-H Bond	sp ³ -sp ³	sp ² -sp ²	sp ³ -sp ²
Length	1.09 Å	1.08 Å	1.08 Å
Strength	413 kJ/mol	439 kJ/mol	410 kJ/mol

Higher s character = stronger and shorter bond

Examples



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