



Study Guide Unit 1-2: *Matter and Measurements*

1. Classify each of the following as an element, compound, homogeneous or heterogeneous mixture.

- (a) air _____ (b) tomato juice _____
(c) iodine (I_2) crystals _____ (d) sand (SiO_2) _____

2. (a) Is a mixture of sugar and sand homogeneous or heterogeneous? _____
(b) Explain how you determined the answer to part (a) _____

3. In 1807 the English chemist Humphry Davy passed an electric current through molten potassium hydroxide and isolated a bright, shiny, conductive substance.

- a. Would you classify the substance as a metal, nonmetal or metalloid? _____
b. Was the change that occurred chemical or physical? Explain. _____

4. Read the following description of the element magnesium. Circle the physical properties and underline the chemical properties.

Magnesium is a silver-gray colored metal with a melt point of 650°C . When magnesium is placed in acid, the metal dissolves releasing hydrogen gas. Magnesium has a hardness on the Mohs scale of 2.5 and a density of 1.74 g/cm^3 at 25°C . It reacts with oxygen gas when heated to form zinc oxide, MgO .

5. A match is lit and held under a cold piece of metal. The following observations are made:

(a) The match burns. (b) The metal gets warmer. (c) Water condenses on the metal. (d) Soot (carbon) is deposited on the metal.

Which letters exhibit occurrences of physical changes? _____ Chemical changes? _____

6. Determine if A-E are observations or interpretations. A match is lit and held under a cold piece of metal. The following occurs:

- a. The match burns. _____ b. The match reacts with oxygen. _____
c. The metal gets warmer. _____ d. Heat transfers to the metal from the match. _____
e. Black soot is deposited on the metal. _____

7. Circle the larger unit: a. centigram or milligram b. liter or kiloliter c. millisecond or microsecond

8. Fill in the blank (Know the following):

- | | | |
|---|--|---|
| a. $1\text{ cm} = \underline{\hspace{1cm}}\text{ mm}$ | c. $1\text{ L} = \underline{\hspace{1cm}}\text{ mL}$ | i. $1\text{ m} = \underline{\hspace{1cm}}\text{ km}$ |
| b. $1\text{ km} = \underline{\hspace{1cm}}\text{ m}$ | e. $1\text{ kg} = \underline{\hspace{1cm}}\text{ g}$ | j. $1\text{ hr} = \underline{\hspace{1cm}}\text{ min}$ |
| c. $1\text{ m} = \underline{\hspace{1cm}}\text{ mm}$ | g. $1\text{ g} = \underline{\hspace{1cm}}\text{ mg}$ | k. $1\text{ min} = \underline{\hspace{1cm}}\text{ sec}$ |
| d. $1\text{ m} = \underline{\hspace{1cm}}\text{ cm}$ | h. $1\text{ g} = \underline{\hspace{1cm}}\text{ ng}$ | |

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James E. Brady, John R. Holum



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Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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Chemistry, Study Guide James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

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Robinson Chemistry Study Guide Robinson, 1992

Chemistry: The Study of Matter Henry Dorin, 1982

Study Guide to Chemistry and Life John William Hill, Dorothy M. Feigl, 1978

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