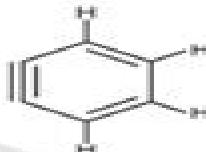


Chemical Bonding

1. How many sp^2 and sp -hybridised carbon atoms are present respectively in the following compound?



(A) 4, 2
(C) 3, 3

2. Give the correct order of initials T or F for following statements. Use T if statement is true and F if it is false:

(i) The order of repulsion between different pair of electrons is $l_p - l_p > l_p - b_p > b_p - b_p$
(ii) In general, as the number of lone pair of electrons on central atom increases, value of bond angle from normal bond angle also increases
(iii) The number of lone pair on O in H_2O is 2 while on N in NH_3 is 1
(iv) The structures of xenon fluorides and xenon oxyfluorides could not be explained on the basis of VSEPR theory

(A) T T T F
(C) T F T T

(B) T F T F
(D) T F F F

3. Which of the following overlaps is incorrect (assuming Z-axis is internuclear axis)?

(i) $2p_y + 2p_y \rightarrow \pi$ -Bond formation
(ii) $2p_x + 2p_x \rightarrow \sigma$ -Bond formation
(iii) $3d_{xy} + 3d_{xy} \rightarrow \pi$ -Bond formation
(iv) $2s + 2p_y \rightarrow \pi$ -Bond formation
(v) $3d_{xy} + 3d_{xy} \rightarrow \delta$ -Bond formation
(vi) $2p_z + 2p_z \rightarrow \sigma$ -Bond formation

(A) i, ii, iii
(C) ii, v

(B) iii, vi
(D) ii, iii, iv

4. Choose the correct code of characteristics for the given order of hybrid orbitals of same atom, $sp < sp^2 < sp^3$

(i) Electronegativity
(ii) Bond angle between same hybrid orbitals

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