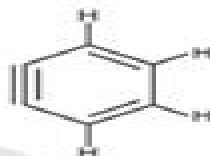


# Chemical Bonding

1. How many  $sp^2$  and  $sp$ -hybridised carbon atoms are present respectively in the following compound?



- (A) 4, 2  
(B) 6, 0  
(C) 3, 3  
(D) 5, 1
2. Give the correct order of initials T or F for following statements. Use T if statement is true and F if it is false:
- The order of repulsion between different pair of electrons is  $l_p - l_p > l_p - b_p > b_p - b_p$
  - In general, as the number of lone pair of electrons on central atom increases, value of bond angle from normal bond angle also increases
  - The number of lone pair on O in  $H_2O$  is 2 while on N in  $NH_3$  is 1
  - The structures of xenon fluorides and xenon oxyfluorides could not be explained on the basis of VSEPR theory
- (A) T T T F  
(B) T F T F  
(C) T F T T  
(D) T F F F
3. Which of the following overlaps is incorrect (assuming Z-axis is internuclear axis)?
- $2p_y + 2p_y \rightarrow \pi$ -Bond formation
  - $2p_x + 2p_x \rightarrow \sigma$ -Bond formation
  - $3d_{xy} + 3d_{xy} \rightarrow \pi$ -Bond formation
  - $2s + 2p_y \rightarrow \pi$ -Bond formation
  - $3d_{xy} + 3d_{xy} \rightarrow \delta$ -Bond formation
  - $2p_z + 2p_z \rightarrow \sigma$ -Bond formation
- (A) i, ii, iii  
(B) iii, vi  
(C) ii, v  
(D) ii, iii, iv
4. Choose the correct code of characteristics for the given order of hybrid orbitals of same atom,  $sp < sp^2 < sp^3$
- Electronegativity
  - Bond angle between same hybrid orbitals

# Pdf Question With Answers On Chemical Bonding

**D Siedentop**



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