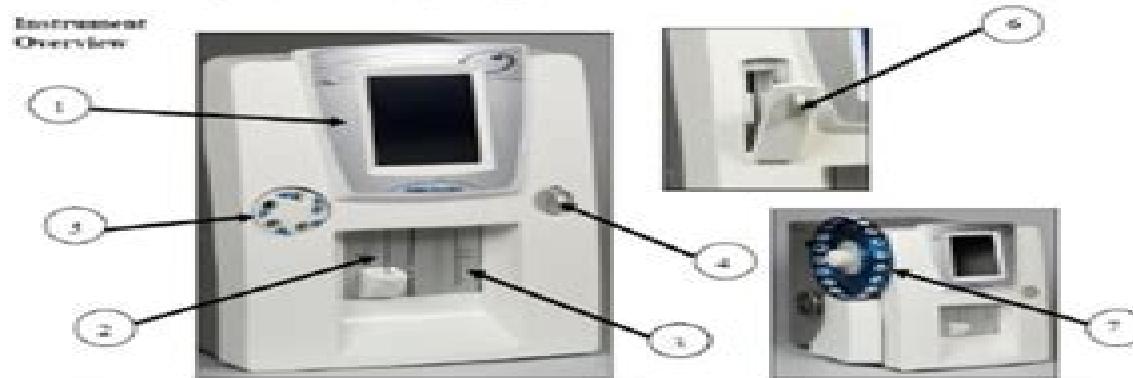




Medonic M-Series Maintenance

The intent of this document is to supplement the information already contained in the Medonic M-Series User's Manual (Section 8) and the Boule Cleaning Kit (1504111) package insert.



Part	Function
1. Display	LCD touch screen with incorporated keyboard and numerical pad.
2. Open Tube (OT) sample probe	Aspirates whole blood from open sample tube.
3. Pre-dilute (PD) probe/ Dispenser	Aspirates pre-diluted samples and dispenses diluent. Used in monthly cleaning procedure.
4. Micropipette Adapter (MPA/MC)	Micro Pipette Adapter enables users to analyze 20 μ L of blood.
5. Mixer (optional)	Uniformly mixes samples.
6. Cap Pierce (CP) (optional)	Analyzes samples with decreased risk of blood contact.
7. Autoloader (AL) (optional)	Enables consecutive samples to be analyzed automatically.

Daily Maintenance

1. Clean the **Open Tube** sample probe and **Pre-dilute** probe using an alcohol wipe.
2. Remove possible traces of salt crystals or blood on the **Open Tube** probe, probe rinse cup, **Pre-dilute** probe, and around the top of the **Cap Pierce** or **Autoloader** using a paper tissue with a disinfecting solution.
3. Clean outside of the instrument with soft cloth and DI water if needed.

Monthly Cleaning Procedure (~10 minutes)

1. Fill a cup with 10 mL 2% hypochlorite from Boule Cleaning Kit (White Cap).
2. Fill another cup with 18 mL of diluent. The diluent can be obtained by using the dispense function on the M-Series. To dispense diluent:
 - a. Press [Dispense] from Main Menu.
 - b. Hold empty cup under **Pre-dilute** probe (1:200 probe on right side).
 - c. Press 1:200 Start plate behind probe
 - d. Discard first dispense.
 - e. Repeat steps b-c four times to achieve 18 mL of diluent.
 - f. Press [Cancel] to exit the dispense function.
3. Aspirate the hypochlorite (White Cap) as a **Pre-dilute** sample (1:200 probe) **twice**.
• Note: *Pre-dilute mode has intentional delay. Press and hold start plate to start cycle (keep probe submerged in hypochlorite solution until the screen displays "Now Analyzing").*

Medonic M Series Service Manual

Lauren Gardner

Medonic M Series Service Manual:

Service Manual for Morris Eight, Series "E", 1939 to 1948, Morris Ten, Series "M", 1939 to 1949, 1956 The Morris Ten ,1960 **MG Service Manual** ,1952 Service manual for Morris Minor Series MM (1949-1953) and II (1953 to 1956). ,1960 Morris Ten (series "M") Workshop Manual Morris Motors, Ltd.. Cowley, England,1954 **VE Series Valiant Service Manual** Chrysler Australia. Service Division,1967 Service Manual for Morris Minor ,1958 **Valiant** 1970-1971 Ellery Publications,1970 **Chrysler Sigma GE Series Service Manual: Part no. 4061761** ,1977 Morris Minor, Series MM, Series 11, and 1100, Workshop Manual British Leyland (Austin-Morris) Limited. Service Division,1970 **Technical Service Manual [for] Hornet, 6 and V-8, 01 Series, Matador, 6 and V-8, 10 Series, Gremlin, 40 Series, Javelin, 6 and V-8, 70 Series, Ambassador, 6 and V-8, 80 Series** American Motors Corporation. Automotive Technical Service,1971 Service Manual for Morris Minor Series MM 1949-1953 (SMM/501 to SMM/177764), Series II 1953-1956 (SMM/177765 Onwards). Scientific Magazines Publishing Co,1960 Service Manual for Morris Minor ,1960 **Service Manual for Morris Minor** ,1955 **GH Series Sigma Service Manual** ,1982 Workshop Manual Minx Series I to V, Husky Series I to III, Cob Series I to III. Rootes Limited. Service Division,1965 **Multicare Series Models 302 & 304** Rigel,198? ... **Oldsmobile Service Manual** General Motors Corporation. Oldsmobile Division,19?? **Technical Service Manual** Nash motors,1954 **An vaterländischen Festtagen und zur Begrüssung von Behörden** ,1890

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Table of Contents Medonic M Series Service Manual

1. Understanding the eBook Medonic M Series Service Manual
 - The Rise of Digital Reading Medonic M Series Service Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Medonic M Series Service Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Medonic M Series Service Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Medonic M Series Service Manual

- Personalized Recommendations
- Medonic M Series Service Manual User Reviews and Ratings
- Medonic M Series Service Manual and Bestseller Lists

5. Accessing Medonic M Series Service Manual Free and Paid eBooks

- Medonic M Series Service Manual Public Domain eBooks
- Medonic M Series Service Manual eBook Subscription Services
- Medonic M Series Service Manual Budget-Friendly Options

6. Navigating Medonic M Series Service Manual eBook Formats

- ePub, PDF, MOBI, and More
- Medonic M Series Service Manual Compatibility with Devices
- Medonic M Series Service Manual Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Medonic M Series Service Manual
- Highlighting and Note-Taking Medonic M Series Service Manual
- Interactive Elements Medonic M Series Service Manual

8. Staying Engaged with Medonic M Series Service Manual

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Medonic M Series Service Manual

9. Balancing eBooks and Physical Books Medonic M Series Service Manual

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Medonic M Series Service Manual

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Medonic M Series Service Manual

- Setting Reading Goals Medonic M Series Service Manual
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Medonic M Series Service Manual

- Fact-Checking eBook Content of Medonic M Series Service Manual
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0 6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no2 which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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process that allows for intermediate chemical structures to be used as precursors to pharmaceuticals and other industrially produced chemicals by adding different functional groups into an aromatic ring

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lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with either a mixture of sulfuric acid or nitric acid creates what is known as a nitro group