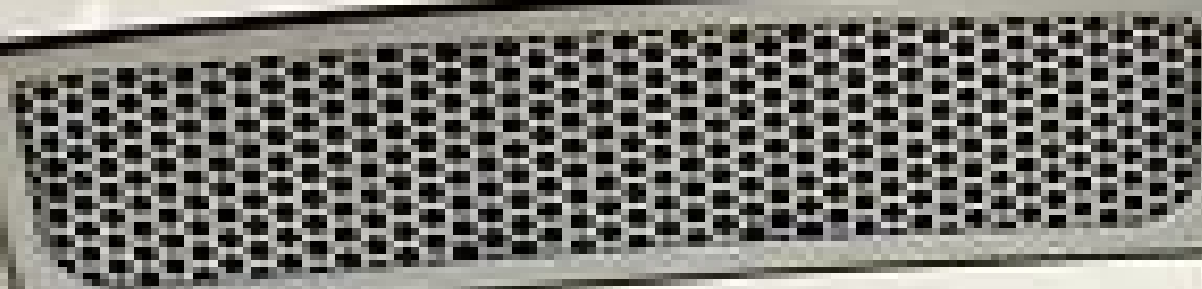


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web what is the mole fraction of liCl in the solution 11.8 practice problem acetone C<sub>3</sub>H<sub>6</sub>O has a density of 0.791 g/ml and xylene C<sub>8</sub>H<sub>10</sub> has a density of 0.864 g/ml a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if 19.0 ml of acetone was dissolved in 98.4 ml of xylene 7.1

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web since the mole fraction of the O<sub>2</sub> is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O<sub>2</sub> are mixed in a flask at room temperature