

MCAT Chemistry – Medical College Admission Test Preparation

→ AROMATIC COMPOUNDS

CHAPTER # 24

AdamjeeCoaching-Blogspot.com

1. Benzene is soluble in

(a) alcohol	(b) petrol
(c) ether	(d) ✓ all of these

2. Benzene is insoluble in

(a) ✓ water	(b) alcohol
(c) petrol	(d) ether

3. Chlorination of benzene is:

(a) ✓ An electrophilic substitution reaction	(b) A nucleophilic substitution reaction
(c) A free-radical reaction	(d) An elimination reaction

4. The total number of sigma bonds in benzene are:

(a) 6	(b) ✓ 12
(c) 18	(d) 3

5. The Pi (π) molecular orbital in benzene contains:

(a) 2 electrons	(b) 4 electrons
(c) ✓ 6 electrons	(d) 3 electrons

6. The C – C – C bond angles in benzene are:

(a) ✓ 120°	(b) 90°
(c) 109.28°	(d) 180°

7. The carbon–carbon bond length in benzene is:

(a) longer than a C – C single bond	(b) ✓ longer than a C = C double bond
(c) Shorter than a C = C double bond	(d) Shorter than a C O \equiv C triple bond

8. Two adjacent substituents on benzene ring is said to be:

(a) ✓ Ortho to one another	(b) Meta to one another
(c) Para to one another	(d) None of the above

Mcqs In Aromatic Compounds

Sabine Zange

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Aromatic Character and Aromaticity G. M. Badger, 1969-01-02 Originally published in 1969 this book analyses the special properties of benzene and its derivatives which give them their distinct aromatic character. The text first discusses the structure of benzene including the light shed on it by physical methods and wave mechanics and the structure of polycyclic benzenoid hydrocarbons and heterocyclic aromatic compounds. Aromaticity is defined as a function of the electronic structure of the molecule or ion in question and it is shown that this has important consequences in respect of bond lengths resonance energies electronic absorption spectra and induced ring currents. Final chapters discuss how far aromatic character and aromaticity are associated with non benzenoid hydrocarbons and with more complex systems.

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