

10.1**THE NATURE OF GASES**
SECTION REVIEW**Objectives**

- Describe the motion of gas particles according to the kinetic theory
- Interpret gas processes by means of kinetic theory

Key Terms

• molecular motion	• gas pressure	• ideal gas law
• kinetic theory	• Boyle's law	• law of combined gases
• general PV =	• Avogadro's law	• law of partial pressures

Part A: Completion

Use this completion section to check your understanding of the concepts in this section. If you need more room for extra practice, please turn over the next page(s) which are blank, otherwise, continue here.

The kinetic theory describes the _____ 1 _____ of gas particles for the motion and the forces of attraction between them. The theory assumes that the volume occupied by a gas is nearly _____ 2 _____, that the particles of gas are _____ 3 _____, and are _____ 4 _____ of each other, and are in constant _____ 5 _____ motion. Furthermore, the _____ 6 _____ between particles are much smaller than the total _____ 7 _____ volume of the gas particles with an object. That is to say, we need to measure _____ 8 _____ pressure. That said, we witness two different types of temperature: _____ 9 _____ and _____ 10 _____.

Part B: True-False

Classify each of these statements as true or false (T) or false (F). **DO NOT** write in the box. **DO** write in the box.

_____ 11. Atmospheric pressure is 2000 times 14.7.

_____ 12. 14.7 is the full gas pressure in the atmosphere.

_____ 13. Atmospheric vapor pressure can be used to convert Celsius to Fahrenheit temperature. That said, 32° Celsius is equivalent to 89.6° Fahrenheit.

Nature Of Gases Section Review Key

M Carnoy

Nature Of Gases Section Review Key:

Reviewing **Nature Of Gases Section Review Key**: Unlocking the Spellbinding Force of Linguistics

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