

15. Test your understanding of isotopes by examining the relationships between the pairs of atoms listed below:

Atom 1	Atom 2	Relationship between atom 1 and atom 2
$^{12}_6\text{C}$	$^{13}_6\text{C}$	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element
Carbon-12	$^{12}_6\text{C}$	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element
Argon-40	Argon-41	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element
$^{11}_5\text{B}$	Boron-10	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element
An atom with 13 protons and 13 neutrons	An atom with 14 protons and 13 neutrons	<input type="checkbox"/> Isotopes <input type="checkbox"/> Same Atom, Not Isotopes of Each Other <input type="checkbox"/> Different Element

EXERCISES

16. The periodic table has a great deal of information about every atom. Using your periodic table, answer the following questions:

- What is the atomic number of chlorine (Cl)? 17
- What is the atomic number of tungsten (W)? 74
- How many protons are there in any Cl atom? 17
- How many protons are there in any Te atom? 52
- Can you tell from the periodic table exactly how many neutrons are in an atom?

No, the mass number is an average mass of all of the isotopes of that element. Since isotopes are based on the number of neutrons, the mass of each isotope of an element will be a little different.

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L Cohen



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