

Chapter 1

Problem Solutions

1.1

- (a) fcc: 8 corner atoms $\times 1/8 = 1$ atom
 6 face atoms $\times 1/2 = 3$ atoms
 Total of 4 atoms per unit cell

- (b) bcc: 8 corner atoms $\times 1/8 = 1$ atom
 1 enclosed atom $= 1$ atom
 Total of 2 atoms per unit cell

- (c) Diamond: 8 corner atoms $\times 1/8 = 1$ atom
 6 face atoms $\times 1/2 = 3$ atoms
 4 enclosed atoms $= 4$ atoms
 Total of 8 atoms per unit cell

1.2

- (a) 4 Ge atoms per unit cell

$$\text{Density} = \frac{4}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\underline{\text{Density of Ge} = 2.22 \times 10^{22} \text{ cm}^{-3}}$$

4 As atoms per unit cell, so that

$$\underline{\text{Density of As} = 2.22 \times 10^{22} \text{ cm}^{-3}}$$

- (b)

8 Ge atoms per unit cell

$$\text{Density} = \frac{8}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\underline{\text{Density of Ge} = 4.44 \times 10^{22} \text{ cm}^{-3}}$$

1.3

- (a) Simple cubic lattice; $a = 2r$

$$\text{Unit cell vol} = a^3 = (2r)^3 = 8r^3$$

$$1 \text{ atom per cell, so atom vol.} = (1) \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{\left(\frac{4\pi r^3}{3} \right)}{8r^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 52.4\%}$$

- (b) Face-centered cubic lattice

$$d = 4r = a\sqrt{2} \Rightarrow a = \frac{d}{\sqrt{2}} = 2\sqrt{2}r$$

$$\text{Unit cell vol} = a^3 = (2\sqrt{2}r)^3 = 16\sqrt{2}r^3$$

$$4 \text{ atoms per cell, so atom vol.} = 4 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{4 \left(\frac{4\pi r^3}{3} \right)}{16\sqrt{2}r^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 74\%}$$

- (c) Body-centered cubic lattice

$$d = 4r = a\sqrt{3} \Rightarrow a = \frac{4}{\sqrt{3}}r$$

$$\text{Unit cell vol.} = a^3 = \left(\frac{4}{\sqrt{3}}r \right)^3$$

$$2 \text{ atoms per cell, so atom vol.} = 2 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{2 \left(\frac{4\pi r^3}{3} \right)}{\left(\frac{4}{\sqrt{3}}r \right)^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 68\%}$$

- (d) Diamond lattice

$$\text{Body diagonal} = d = 8r = a\sqrt{3} \Rightarrow a = \frac{8}{\sqrt{3}}r$$

$$\text{Unit cell vol.} = a^3 = \left(\frac{8r}{\sqrt{3}} \right)^3$$

$$8 \text{ atoms per cell, so atom vol.} = 8 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{8 \left(\frac{4\pi r^3}{3} \right)}{\left(\frac{8r}{\sqrt{3}} \right)^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 34\%}$$

1.4

From Problem 1.3, percent volume of fcc atoms is 74%. Therefore after coffee is ground,

$$\underline{\text{Volume} = 0.74 \text{ cm}^3}$$

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web social security and welfare programs and highlights its unique method of financing social security by using the mandatory national provident fund mechanism known as fical central provident fund fl or cpf concluding remarks are in the final section which also identifies some file lessons fl for other developing countries

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web social supremacy blueprint by greg greenway is designed for guys who want the ability to meet and date many quality women through being introduced to them via their social circles it s about building a lifestyle that continually adds quality women and guys over time so that you re creating something that offers long term value

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