

13. These calculations have used  $L = 6.02 \times 10^{23}$   
 (a)  $7.22 \times 10^{23}$  (b)  $3.01 \times 10^{23}$

14. 0.53 mol H

15. 0.250 mol

16. (a)  $262.87 \text{ g mol}^{-1}$  (b)  $176.14 \text{ g mol}^{-1}$   
 (c)  $184.10 \text{ g mol}^{-1}$  (d)  $248.22 \text{ g mol}^{-1}$

17. 189.1 g

18. 1.5 mol

19. 0.0074 mol  $\text{Cl}^-$

20.  $1.83 \times 10^{23}$  C atoms

21. 17.1 g (integer value because no calculator)

22. 10.0 g  $\text{H}_2\text{O}$

23.  $2.0 \text{ mol N}_2 > 3.0 \text{ mol NH}_3 > 25.0 \text{ mol H}_2 > 1.0 \text{ mol N}_2\text{H}_4$

24. (a)  $\text{CH}$  (b)  $\text{CH}_2\text{O}$   
 (c)  $\text{C}_2\text{H}_4\text{O}_2$  (d)  $\text{C}_2\text{H}_6$   
 (e)  $\text{C}_2\text{H}_4$  (f)  $\text{CH}_2\text{O}$

25.  $\text{Na}_2\text{S}_2\text{O}_3$

26.  $\text{CoSO}_4 \cdot 7\text{H}_2\text{O}$

27.  $\text{C}_2\text{H}_5\text{N}$

28.  $\text{NH}_3$

29. 6.94 L

30.  $\text{CoS}$

31. empirical formula  $\text{CH}$ ; molecular formula  $\text{C}_2\text{H}_6$

32. empirical formula  $\text{H}_2\text{PO}_4^-$ ; molecular formula  $\text{H}_4\text{P}_2\text{O}_7$

33.  $\text{C}_{12}\text{H}_{16}\text{N}_4\text{P}_2\text{O}_{10}$ , for both empirical and molecular formulas

34.  $\text{C}_2\text{H}_4\text{O}$

35. Let  $y = \text{mass of chalk in grams}.$   
 $\text{moles of chalk used} = \frac{\text{mass used}}{M(\text{CaCO}_3)} \times$   
 $y \frac{\text{g}}{100.09 \text{ g mol}^{-1}}$

36. This is the same as the number of moles of carbon atoms used.  
 Therefore the number of carbon atoms used =  
 $\text{moles of chalk} \times (6.02 \times 10^{23} \text{ mol}^{-1}) =$   
 $6.02 \times 10^{23} y$   
 $100.09$

37. (a)  $2\text{C}_2\text{H}_6 + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$   
 (b) 1.59 g

38. 4.355 kg

39. (a)  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$   
 (b) 92.8%  
 (c)  $\text{CaCO}_3$  is the only source of  $\text{CO}_2$ ; all the  $\text{CaCO}_3$  undergoes complete decomposition; all  $\text{CO}_2$  released is captured; heating does not cause any change in the mass of the other minerals present.

40. (a) 65.2 g (b) 1.3 g  $\text{H}_2$

41. 5.23 g  $\text{C}_2\text{H}_4\text{O}_2$

42. 253.2 g theoretical  $\text{CaSO}_4$ ; 77.9%

43. 3.16 g ester

44. 107 g of  $\text{C}_2\text{H}_6$  needed

45. (a) 2.40 mol (b) 0.0110 mol

46. (a) 35.85  $\text{dm}^3$  (b) 5.7  $\text{dm}^3$

47. 0.652  $\text{dm}^3$

48. 0.138 mol  $\text{Br}_2$  and 0.156 mol  $\text{Cl}_2$ , so more molecules of  $\text{Cl}_2$

49. 0.113  $\text{dm}^3$

50. 0.28  $\text{dm}^3$

51. 90 kPa

52. 16 °C

53. 3.0  $\text{dm}^3$

54. 2.8  $\text{dm}^3$

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## **Table of Contents Pearson Chemistry Workbook Answers Chapter 2**

1. Understanding the eBook Pearson Chemistry Workbook Answers Chapter 2
  - The Rise of Digital Reading Pearson Chemistry Workbook Answers Chapter 2
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Workbook Answers Chapter 2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chemistry Workbook Answers Chapter 2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Workbook Answers Chapter 2
  - Personalized Recommendations
  - Pearson Chemistry Workbook Answers Chapter 2 User Reviews and Ratings
  - Pearson Chemistry Workbook Answers Chapter 2 and Bestseller Lists
5. Accessing Pearson Chemistry Workbook Answers Chapter 2 Free and Paid eBooks
  - Pearson Chemistry Workbook Answers Chapter 2 Public Domain eBooks
  - Pearson Chemistry Workbook Answers Chapter 2 eBook Subscription Services
  - Pearson Chemistry Workbook Answers Chapter 2 Budget-Friendly Options

6. Navigating Pearson Chemistry Workbook Answers Chapter 2 eBook Formats
  - ePUB, PDF, MOBI, and More
  - Pearson Chemistry Workbook Answers Chapter 2 Compatibility with Devices
  - Pearson Chemistry Workbook Answers Chapter 2 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson Chemistry Workbook Answers Chapter 2
  - Highlighting and Note-Taking Pearson Chemistry Workbook Answers Chapter 2
  - Interactive Elements Pearson Chemistry Workbook Answers Chapter 2
8. Staying Engaged with Pearson Chemistry Workbook Answers Chapter 2
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Chemistry Workbook Answers Chapter 2
9. Balancing eBooks and Physical Books Pearson Chemistry Workbook Answers Chapter 2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Chemistry Workbook Answers Chapter 2
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Workbook Answers Chapter 2
  - Setting Reading Goals Pearson Chemistry Workbook Answers Chapter 2
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Workbook Answers Chapter 2
  - Fact-Checking eBook Content of Pearson Chemistry Workbook Answers Chapter 2
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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