

Chapter 10 test**Matching***Match each item with the correct statement below.*

a. representative particle	d. percent composition
b. mole	e. standard temperature and pressure
c. Avogadro's number	f. empirical formula

_____ 1. the number of representative particles of a substance present in 1 mole of that substance

_____ 2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists

_____ 3. the SI unit used to measure amount of substance

_____ 4. 0°C and 1 atm

_____ 5. the percent by mass of each element in a compound

_____ 6. the smallest whole number ratio of the atoms in a compound

Multiple Choice*Identify the letter of the choice that best completes the statement or answers the question.*

_____ 7. What SI unit is used to measure the number of representative particles in a substance?

a. kilogram	c. kelvin
b. ampere	d. mole

_____ 8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C_3H_7OH ?

a. $5 \times (6.02 \times 10^{23})$	c. 35
b. 5	d. $35 \times (6.02 \times 10^{23})$

_____ 9. All of the following are equal to Avogadro's number EXCEPT _____.

a. the number of atoms of bromine in 1 mol Br_2
b. the number of atoms of gold in 1 mol Au
c. the number of molecules of nitrogen in 1 mol N_2
d. the number of molecules of carbon monoxide in 1 mol CO

_____ 10. How many moles of tungsten atoms are in 4.8×10^{23} atoms of tungsten?

a. 8.0×10^{-2} moles	c. 1.3×10^{-4} moles
b. 8.0×10^1 moles	d. 1.3×10^{-2} moles

_____ 11. How many atoms are in 0.075 mol of titanium?

a. 1.2×10^{23}	c. 6.4×10^2
b. 2.2×10^{24}	d. 4.5×10^{22}

_____ 12. How many molecules are in 2.10 mol CO_2 ?

a. 2.53×10^{24} molecules	c. 3.49×10^{24} molecules
b. 3.79×10^{24} molecules	d. 1.26×10^{24} molecules

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