

13 These calculations have used  $L = 6.02 \times 10^{23}$

- (a)  $7.22 \times 10^{22}$  (b)  $3.01 \times 10^{23}$   
 (c)  $1.20 \times 10^{23}$

14 0.53 mol H

15 0.250 mol

- 16 (a) 262.87 g mol<sup>-1</sup> (b) 176.14 g mol<sup>-1</sup>  
 (c) 184.10 g mol<sup>-1</sup> (d) 248.22 g mol<sup>-1</sup>

17 189.1 g

18 1.5 mol

19 0.0074 mol Cl<sup>-</sup>

20  $1.83 \times 10^{23}$  C atoms

21 171 g (integer value because no calculator)

22 10.0 g H<sub>2</sub>O

23 2.0 mol N<sub>2</sub> > 3.0 mol NH<sub>3</sub> > 25.0 mol H<sub>2</sub> >  
 1.0 mol N<sub>2</sub>H<sub>4</sub>

- 24 (a) CH (b) CH<sub>2</sub>O  
 (c) C<sub>2</sub>H<sub>2</sub>O<sub>2</sub> (d) C<sub>2</sub>H<sub>4</sub>  
 (e) C<sub>2</sub>H<sub>4</sub> (f) CH<sub>2</sub>O

25 Na<sub>2</sub>S<sub>2</sub>O<sub>5</sub>

26 CoSO<sub>4</sub>·7H<sub>2</sub>O

27 C<sub>10</sub>H<sub>16</sub>N

28 NH<sub>3</sub>

29 5.94 L

30 CuS

31 empirical formula CH; molecular formula C<sub>7</sub>H<sub>8</sub>

32 empirical formula H<sub>3</sub>PO<sub>3</sub>; molecular formula H<sub>4</sub>P<sub>2</sub>O<sub>6</sub>

33 C<sub>12</sub>H<sub>12</sub>N<sub>2</sub>P<sub>2</sub>O<sub>12</sub> for both empirical and molecular formulas

34 C<sub>7</sub>H<sub>8</sub>O

35 Let  $y$  = mass of chalk in grams.

$$\text{moles of chalk used} = \frac{\text{mass used}}{M_r(\text{CaCO}_3)} \times$$

$$y \frac{1 \text{ g}}{100.09 \text{ g mol}^{-1}}$$

This is the same as the number of moles of carbon atoms used.

$$\begin{aligned} \text{Therefore the number of carbon atoms used} &= \\ \text{moles of chalk} \times (6.02 \times 10^{23} \text{ mol}^{-1}) &= \\ \frac{6.02 \times 10^{23} y}{100.09} \end{aligned}$$

- 36 (a) 2.50 mol (b) 5.63 mol  
 (c) 665.5 g

37 (a)  $2\text{C}_2\text{H}_6 + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$   
 (b) 1.59 g

38 4.355 kg

- 39 (a)  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$   
 (b) 92.8%

(c) CaCO<sub>3</sub> is the only source of CO<sub>2</sub>; all the CaCO<sub>3</sub> undergoes complete decomposition; all CO<sub>2</sub> released is captured; heating does not cause any change in the mass of the other minerals present.

- 40 (a) 65.2 g (b) 1.3 g H<sub>2</sub>

41 5.23 g C<sub>7</sub>H<sub>8</sub>Cl<sub>2</sub>

42 253.2 g theoretical CaSO<sub>4</sub>; 77.9%

43 3.16 g ester

44 107 g of C<sub>2</sub>H<sub>6</sub> needed

- 45 (a) 2.40 mol (b) 0.0110 mol  
 (c) 44 mol

- 46 (a) 35.65 dm<sup>3</sup> (b) 5.7 dm<sup>3</sup>

47 0.652 dm<sup>3</sup>

48 0.138 mol Br<sub>2</sub> and 0.156 mol Cl<sub>2</sub>, so more molecules of Cl<sub>2</sub>

49 0.113 dm<sup>3</sup>

50 0.28 dm<sup>3</sup>

51 90 kPa

52 16 °C

53 3.0 dm<sup>3</sup>

54 2.8 dm<sup>3</sup>

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### **Table of Contents Pearson Chemistry Study Workbook Answers 4826**

1. Understanding the eBook Pearson Chemistry Study Workbook Answers 4826
  - The Rise of Digital Reading Pearson Chemistry Study Workbook Answers 4826
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Study Workbook Answers 4826
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Chemistry Study Workbook Answers 4826
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Study Workbook Answers 4826
  - Personalized Recommendations
  - Pearson Chemistry Study Workbook Answers 4826 User Reviews and Ratings
  - Pearson Chemistry Study Workbook Answers 4826 and Bestseller Lists

5. Accessing Pearson Chemistry Study Workbook Answers 4826 Free and Paid eBooks
  - Pearson Chemistry Study Workbook Answers 4826 Public Domain eBooks
  - Pearson Chemistry Study Workbook Answers 4826 eBook Subscription Services
  - Pearson Chemistry Study Workbook Answers 4826 Budget-Friendly Options
6. Navigating Pearson Chemistry Study Workbook Answers 4826 eBook Formats
  - ePub, PDF, MOBI, and More
  - Pearson Chemistry Study Workbook Answers 4826 Compatibility with Devices
  - Pearson Chemistry Study Workbook Answers 4826 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson Chemistry Study Workbook Answers 4826
  - Highlighting and Note-Taking Pearson Chemistry Study Workbook Answers 4826
  - Interactive Elements Pearson Chemistry Study Workbook Answers 4826
8. Staying Engaged with Pearson Chemistry Study Workbook Answers 4826
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Chemistry Study Workbook Answers 4826
9. Balancing eBooks and Physical Books Pearson Chemistry Study Workbook Answers 4826
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Chemistry Study Workbook Answers 4826
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Pearson Chemistry Study Workbook Answers 4826
  - Setting Reading Goals Pearson Chemistry Study Workbook Answers 4826
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Chemistry Study Workbook Answers 4826
  - Fact-Checking eBook Content of Pearson Chemistry Study Workbook Answers 4826
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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