

1. States of matter

SC1.1w States of matter

- 1 Solid: particles are close together; particles are regularly arranged; particles can only vibrate about fixed positions; diagram chosen matches the written descriptions.
Liquid: particles are close together; particles are randomly arranged; particles can move around each other; diagram chosen matches the written descriptions.
Gas: particles are far apart; particles are randomly arranged; particles can move quickly in all directions; diagram chosen matches the written descriptions.
- 2 Student compares the arrangements in each state – regularly arranged in solids; randomly arranged in gases and liquids. Answer compares the movement in each state – can only vibrate in solids; can move around each other in liquids; can move quickly in all directions in gases

SC1.3w Changing states – Strengthen

- 1 Suitable particle diagrams drawn in three boxes. Gas: particles separate; (at least 10 diameters apart if possible); random arrangement. Liquid: particles touching; random arrangement. Solid: particles touching; regular arrangement
- 2 Descriptions of arrangement and movements of particles in each state, written next to relevant particle diagrams. Gas arrangement: random; far apart. Gas movement: rapid in all directions. Liquid arrangement: random; close together. Liquid movement: can move around each other. Solid arrangement: regular; close together. Solid movement: can only vibrate about fixed positions
- 3 Correct state change names written in three boxes; with descriptions of changes to arrangement and movement of particles included.
Energy at top: high; energy at bottom: low
- 4 Should include the idea that the temperature will rise above 0 °C (melting point of water) in the spring or summer; so the ice hotel will melt.

SC1.4w Changing states – Extend

- 1 A Energy is transferred from the surroundings, to the particles.
B Energy is transferred to the surroundings, from the particles

- 2 A Melting: become able to move around each other; regular to random arrangement. Freezing: can no longer move around each other; random to regular arrangement.
B Boiling: still randomly arranged; but become far apart; and free to move in all directions. Condensing: still randomly arranged; but become close together; and can only move around each other
- 3 A Particles become close together; and regularly arranged (rather than randomly arranged); can no longer move in all directions; but vibrate in fixed positions.
B Particles become far apart; and randomly arranged; rather than regularly arranged; can move in all directions; rather than just vibrate in fixed positions
- 4 A The attractive forces must be weak; if they are easily overcome / formed.
B It is cold and solid; it is dry because it does not melt to form a liquid

SC1.5w Particle theory and states of matter – Homework 1

- 1 a Liquid
b Solid
c Gas
- 2 Particles are still randomly arranged; but become far apart and free to move in all directions.
- 3 No new substances are formed
- 4 a Horizontal part of curve labelled
b 55 °C labelled
c Curve from 0–7 minutes labelled
d Curve at 0 minutes labelled
- 5 Solid at –150 °C; gas at 25 °C
- 6 Its melting point must be below 25 °C; its boiling point must be above 25 °C.

SC1.6w Particle theory and states of matter – Homework 2

- 1 a Solid; liquid; gas
b Melting: solid → liquid; sublimation: solid → gas; boiling/evaporation: liquid → gas; condensing: gas → liquid; deposition: gas → solid; freezing/solidifying: liquid → solid
- 2 a Close together; randomly arranged; can move around each other
b Close together; regularly arranged; can only vibrate about fixed positions

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