

1. States of matter

SC1.1w States of matter

- 1** *Solid*: particles are close together; particles are regularly arranged; particles can only vibrate about fixed positions; diagram chosen matches the written descriptions.
Liquid: particles are close together; particles are randomly arranged; particles can move around each other; diagram chosen matches the written descriptions.
Gas: particles are far apart; particles are randomly arranged; particles can move quickly in all directions; diagram chosen matches the written descriptions.
- 2** Student compares the arrangements in each state – regularly arranged in solids; randomly arranged in gases and liquids. Answer compares the movement in each state – can only vibrate in solids; can move around each other in liquids; can move quickly in all directions in gases.

SC1.3w Changing states – Strengthen

- 1** Suitable particle diagrams drawn in three boxes. *Gas*: particles separate; (at least 10 diameters apart if possible); random arrangement. *Liquid*: particles touching; random arrangement. *Solid*: particles touching; regular arrangement.
- 2** Descriptions of arrangement and movements of particles in each state, written next to relevant particle diagrams. *Gas* arrangement: random; far apart. *Gas* movement: rapid in all directions. *Liquid* arrangement: random; close together. *Liquid* movement: can move around each other. *Solid* arrangement: regular; close together. *Solid* movement: can only vibrate about fixed positions.
- 3** Correct state change names written in three boxes; with descriptions of changes to arrangement and movement of particles included.
Energy at top: high; energy at bottom: low
- 4** Should include the idea that the temperature will rise above 0 °C (melting point of water) in the spring or summer; so the ice hotel will melt.

SC1.4w Changing states – Extend

- 1** **A** Energy is transferred from the surroundings, to the particles
B Energy is transferred to the surroundings, from the particles

- 2** **A** *Melting*: become able to move around each other; regular to random arrangement. *Freezing*: can no longer move around each other; random to regular arrangement
B *Boiling*: still randomly arranged; but become far apart; and free to move in all directions. *Condensing*: still randomly arranged; but become close together; and can only move around each other
- 3** **A** Particles become close together; and regularly arranged (rather than randomly arranged); can no longer move in all directions; but vibrate in fixed positions
B Particles become far apart; and randomly arranged; rather than regularly arranged; can move in all directions; rather than just vibrate in fixed positions
- 4** **A** The attractive forces must be weak; if they are easily overcome / formed
B It is cold and solid; it is dry because it does not melt to form a liquid

SC1.5w Particle theory and states of matter – Homework 1

- 1** **a** Liquid
b Solid
c Gas
- 2** Particles are still randomly arranged; but become far apart and free to move in all directions
- 3** No new substances are formed
- 4** **a** Horizontal part of curve labelled
b 55 °C labelled
c Curve from 0–7 minutes labelled
d Curve at 0 minutes labelled
- 5** Solid at –150 °C; gas at 25 °C
- 6** Its melting point must be below 25 °C; its boiling point must be above 25 °C.

SC1.6w Particle theory and states of matter – Homework 2

- 1** **a** Solid; liquid; gas
b *Melting*: solid → liquid; *sublimation*: solid → gas; *boiling/evaporation*: liquid → gas; *condensing*: gas → liquid; *deposition*: gas → solid; *freezing/solidifying*: liquid → solid
c Solid
- 2** **a** Close together; randomly arranged; can move around each other
b Close together; regularly arranged; can only vibrate about fixed positions

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