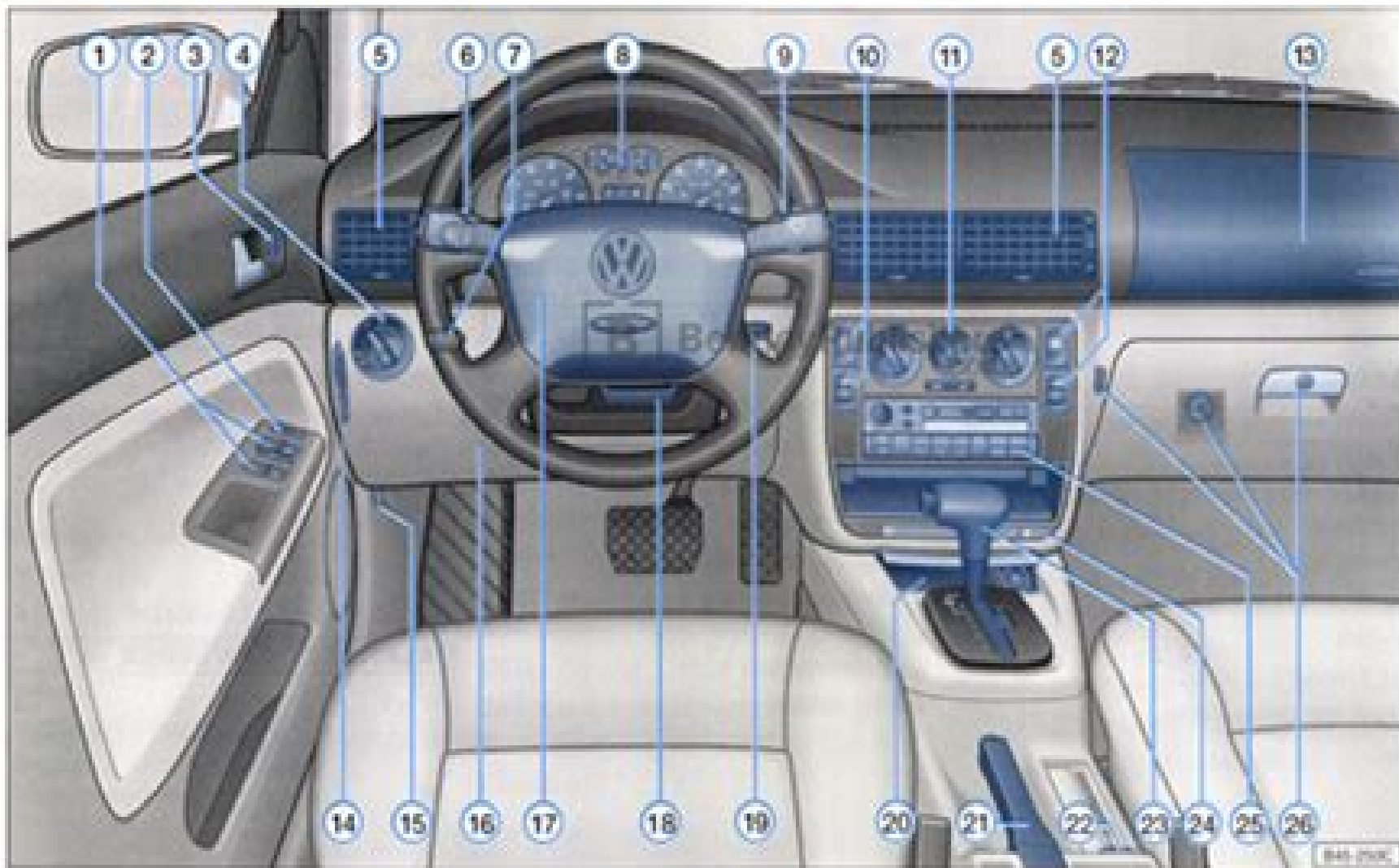


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# Passat 1999 Manual

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## **Passat 1999 Manual:**

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forward reaction rate is now equal to the reverse reaction rate in regards to the amounts or concentrations of the reactants  
and the products there is no change due to the forward reaction rate being equal to the reverse reaction rate 2

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web the rate of appearance of a product or the rate of disappearance of a reactant the change in the amount of substance  
over time reaction mechanism the series of steps by which reactions occur activation energy minimum energy needed to  
cause a reaction to occur activated complex

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11 3 ml min when is the rate of the reaction the greatest slope was steepest 3 4 min time interval when slope

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and products are present some reaction mixtures contain mostly reactants and form only a few products at equilibrium some  
reaction mixtures contain mostly products and remain mostly reactants at equilibrium

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web reaction rate the number of particles that react in a given time to form products le châtelier's principle if a stress is applied to a system in dynamic equilibrium the system changes to relieve the stress chemical equilibrium the forward and reverse reactions take place at the same rate activation energy

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web 1 the change in pressure will only effect gaseous equilibrium 2 increase the pressure will usually the direction that has fewer molecules  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$  for every two molecules of ammonia made four molecules of reactant are used up this equilibrium shifts to the right with an increase in pressure

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web if the rate of the forward reaction decreases increases stays the same more products are formed in a chemical reaction at equilibrium the rate of the forward reaction is equal to the rate of the reverse reaction if the rate of the forward reaction increases more products are formed

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web in order to physically measure the rate of consumption of the bisulfite ion we will look for a color change in the reaction vessel the solution will change from clear to dark blue if you have the following equilibrium reaction  $\text{CO(g)} + 2\text{H}_2\text{(g)} \rightleftharpoons \text{CH}_3\text{OH(g)}$

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web the energy needed for a reaction to occur the minimum amount of energy needed for a reaction to potentially occur the energy added by a catalyst the energy possessed by the products question 5 60 seconds q increase in temperature of the reactants can do one of the following answer choices

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