

1. Complete and classify the following partial equations.

- $\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{Fe}^{2+} + \text{H}_2$ (single replacement) $\text{2Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$
- $\text{Na} + \text{Cl}_2 \rightarrow$
- $2\text{Zn} + \text{Pb(NO}_3\text{)}_2 \rightarrow 2\text{Zn}^{2+}(\text{NO}_3\text{)}_2 + \text{Pb}$ (single replacement)
- $\text{NaNO}_3 + \text{HCl} \rightarrow$
- $2\text{C}_6\text{H}_6 + \text{O}_2 \rightarrow 2\text{C}_6\text{H}_5\text{OH} + \text{H}_2\text{O}$ (combustion)
- $\text{Al}_2\text{O}_3 \rightarrow$
- $\text{N}_2 + \text{H}_2 \rightarrow 2\text{NH}_3$ (synthesis)
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- $2\text{Na} + \text{H}_2\text{O} \rightarrow 2\text{Na}^{+}(\text{H}_2\text{O})_2 + \text{H}_2$ (displacement)
- $\text{Na}_2\text{CO}_3 + \text{H}_3\text{PO}_4 \rightarrow \text{Na}^{+}_2\text{PO}_4 + 3\text{H}_2\text{O}$ (acid-base)
- $\text{H}_2\text{SO}_4 + \text{Al(OH)}_3 \rightarrow$
- $2\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 2\text{C}_2\text{H}_5\text{OH}$ (combustion)
- $\text{NaCl} \rightarrow$
- $\text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{K}^{+}\text{SO}_4 + 2\text{H}_2\text{O}$ (acid-base)
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- $\text{NaClO} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ (acid-base)
- $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$ (combustion)
- $\text{H}_2\text{O} \rightarrow$
- $\text{Fe}_2\text{O}_3 + \text{H}_2\text{CO}_3 \rightarrow \text{Fe}_2(\text{CO}_3)_3 + \text{H}_2$ (displacement)
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- $\text{Pb} + \text{H}_2\text{SO}_4 \rightarrow \text{PbSO}_4 + \text{H}_2$ (displacement)
- $\text{K}^{+} + \text{H}_2\text{O} \rightarrow$
- $\text{Ba}(\text{OH})_2 + \text{H}_2\text{SO}_4 \rightarrow \text{Ba}(\text{SO}_4)_2 + 2\text{H}_2\text{O}$ (acid-base)
- $\text{Pb} + \text{CuSO}_4 \rightarrow$

2. Given: $2\text{NaNO}_3 \rightarrow 2\text{Na} + \text{O}_2$ calculate the mass of oxygen gas produced if 145 g of NaNO_3 is decomposed.

65.4%

3. Calculate the mass of oxygen gas required to burn 240 g of wood ($\text{C}_5\text{H}_8\text{O}_2$).

25.5%

4. If 14.0 moles of butane (C_4H_{10}) is burned:

- calculate the number of moles of oxygen gas required. 49.0 mol O_2
- calculate the number of moles of carbon dioxide produced. 14.0 mol CO_2
- calculate the mass of water produced. $70.4 \text{ g H}_2\text{O}$

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate.

- Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. 2.7 g Al
- Calculate the mass of zinc that would be produced from 50.0 g of aluminum and zinc nitrate. 17.2 g Zn

6. Given: $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$

- How many grams of Cu are required to react with 3.00 g of silver nitrate? 0.56 g Cu
- How many grams of silver are produced from 3.00 g of silver nitrate and sufficient zinc nitrate? 1.50 g Ag



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