

KEY

1. Complete and classify the following partial equations

- $2Fe + 3H_2SO_4 \rightarrow 2FeSO_4 + 3H_2$ (single repl)
- $2Fe + 3H_2SO_4 \rightarrow 2FeSO_4 + 3H_2$ (single repl)
- $12s + 4Pb(NO_3)_2 \rightarrow 12s(PbO_2)_2 + 1Pb$ (single repl)
- $Al_2O_3 + HCl \rightarrow$
- $2Cu + O_2 \rightarrow 2CuO$ $4Cu + 1O_2 \rightarrow 2Cu_2O$ (synthesis)
- $Al_2O_3 \rightarrow$
- $N_2 + 3H_2 \rightarrow 2NH_3$ (synth)
- $C_2H_6 + O_2 \rightarrow$
- $1Mg + 2HNO_3 \rightarrow 1Mg(NO_3)_2 + 1H_2$ (s.r.)
- $Na + H_2O \rightarrow$
- $3NaOH + H_3PO_4 \rightarrow 1Na_3PO_4 + 3H_2O$ (acid base)
- $H_2SO_4 + Al(OH)_3 \rightarrow$
- $2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$ (combustion)
- $AlCl_3 \rightarrow$
- $2KOH + H_2SO_4 \rightarrow 1K_2SO_4 + 2H_2O$ (acid base)
- $C_2H_6 + O_2 \rightarrow$
- $1NaOH + HCl \rightarrow 1NaCl + 1H_2O$ (acid base)
- $NaHCO_3 + H_2SO_4 \rightarrow$
- $1C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$ (combustion)
- $H_2O \rightarrow$
- $2HNO_3 + Na_2CO_3 \rightarrow 1Na_2CO_3 + 1Ca(NO_3)_2$ (s.r.)
- $CH_4 + O_2 \rightarrow$
- $1Pb + 1H_2SO_4 \rightarrow 1PbSO_4 + 1H_2$ (s.r.)
- $KI + Br_2 \rightarrow$
- $1Ba(OH)_2 + 2HNO_3 \rightarrow 1Ba(NO_3)_2 + 2H_2O$ (acid base)
- $Pb + Cu(NO_3)_2 \rightarrow$

2. Given: $2NaClO_3 \rightarrow 2NaCl + 3O_2$ calculate the mass of oxygen gas produced if 143 g of $NaClO_3$ is decomposed

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($C_{21}H_{44}$)

254g

4. If 14.0 moles of butane (C_4H_{10}) is burned

- calculate the number of moles of oxygen gas required 90.0 mol
- calculate the number of moles of carbon dioxide produced 56.0 mol CO_2
- calculate the mass of water produced 108.0 mol H_2O

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate

- Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. 27.4g Al
- Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. 182 g Zn

6. Given: $Cu + 2AgNO_3 \rightarrow Cu(NO_3)_2 + 2Ag$

- How many grams of Cu are required to react with 3.00 g of silver nitrate? 0.544g Cu
- How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? 1.70g Ag

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