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web balancing chemical equations worksheet answer key balance the equations below  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$   $\text{KClO}_3 \rightarrow \text{KCl} + 3\text{O}_2$   $2\text{NaCl} \rightarrow \text{Na}_2\text{Cl}_2$   $\text{Cl}_2 + \text{H}_2 \rightarrow 2\text{HCl}$   $\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$   $\text{Pb(OH)}_2 \rightarrow \text{PbO} + \text{H}_2\text{O}$   $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_2 + \text{HClO}$   $\text{PbCl}_2 + \text{AlBr}_3 \rightarrow \text{AlCl}_3 + \text{PbBr}_2$   $\text{K}_2\text{SO}_4 \rightarrow \text{K}_2\text{O} + \text{SO}_2$   $\text{KBr} \rightarrow \text{K} + \text{Br}_2$   $\text{Al}_2\text{O}_3 + \text{SO}_3 \rightarrow \text{Al}_2\text{SO}_5$   $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

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web may 28 2020 write a balanced chemical equation for each step of the process the first step is the decomposition of solid calcium carbonate from seashells to form solid calcium oxide and gaseous carbon dioxide the second step is the formation of solid calcium hydroxide as the only product from the reaction of the solid calcium oxide with liquid water