REDOX REACTIONS.

NOTES...

Intro

"Many Chemical reactions involve transfer of electrons from one chemical substance to another. These electron-transfer reactions are termed as Oxidation-Reduction or Redox reactions.

Oxidation

(i) Loss of electron

- Example: $Fe^{2+} \rightarrow Fe^{3+} + e^-$

- Sub-points:

  - Loss of hydrogen

  - Gain of oxygen

  - Gain of electronegative element OR loss of electropositive element

Reduction

i) Gain of electron

- Example: $Cl_2 + 2e^- \rightarrow 2Cl^-$

- Sub-points:

 - Gain of hydrogen

 - Loss of oxygen

  - Loss of electronegative element OR gain of electropositive element

Redox reaction

The reaction in which oxidation as well as reduction takes place together is known as a redox reaction.

"REDOX REACTIONS IN TERMS OF ELECTRON TRANSFER (MODERN IDEA)

Oxidation process is the process in which one or more electrons are lost by an atom, ion, or molecule.

Example: $Zn \rightarrow Zn^{2+} + 2e^-$

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Terry C. Jones

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